

# Acids & Bases Study Guide

6. Water is considered neutral. Describe what this means in terms of  $[H^+]$  &  $[OH^-]$  concentrations.

8. Complete the following:

$$pH = -\log[H^+]$$

$$K_w = [H^+][OH^-]$$

7. Fill in the blank:

9. What is the  $[OH^-]$  of a solution with a pH of 10?

10. What is the  $[H^+]$  of a solution with a pH of 4?

11. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-5}$  M?

12. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-4}$  M?

13. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-12}$  M?

14. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-10}$  M?

15. What is the pH of a solution with a  $[H^+]$  of  $1 \times 10^{-11}$  M?

16. What is the pH of a solution with a  $[OH^-]$  of  $1 \times 10^{-13}$  M?

Name: \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_

## Acids & Bases Study Guide

1. Fill in the blanks or complete and contrast the properties of Acids & Bases.

Acids	Bases
<ul style="list-style-type: none"> <li>• Corrosive _____</li> <li>• Tastes _____</li> <li>• Feels _____</li> <li>• Turns litmus paper _____</li> <li>• Reacts with _____</li> </ul>	<ul style="list-style-type: none"> <li>• Corrosive _____</li> <li>• Tastes _____</li> <li>• Feels _____</li> <li>• Turns litmus paper _____</li> <li>• Reacts with _____</li> </ul>

2. Fill in the chart for the following acids & bases.

Acid or Base?	Strong or Weak?	Formula	Name
			hydrochloric acid
			sulfuric acid
		$HNO_3$	
		$NaOH$	
		$Ca(OH)_2$	
			hydrochloric acid

3. Predict the products for the following double replacement reactions involving acids & bases. Identify which component is the Acid, Base, Conjugate Acid (CA), and Conjugate Base (CB) in the formula.



4. Label the pH scale from 0 to 14. Acidic, Basic, Neutral, Strong Acid, Weak Acid, Strong Base, Weak Base.



# Chemistry Study Guide Acid And Bases Answers

**David R. Klein, Laurie S. Starkey**



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