

# Acids & Bases

## Study Guide

1. Write a balanced nuclear equation when this reaction is carried out at 25°C & 1 atm concentration.

2. Calculate the pH of a 0.1 M  $\text{H}_2\text{SO}_4$  solution.

3. pH of dilution

4. Calculate the pH of a 0.1 M  $\text{NaOH}$  solution.

5. Calculate the pH of a 0.1 M  $\text{HCl}$  solution.

6. Calculate the pH of a 0.1 M  $\text{CH}_3\text{COOH}$  solution.

7. Calculate the pH of a 0.1 M  $\text{NH}_3$  solution.

8. Calculate the pH of a 0.1 M  $\text{H}_2\text{O}_2$  solution.

9. Calculate the pH of a 0.1 M  $\text{H}_2\text{S}$  solution.

10. Calculate the pH of a 0.1 M  $\text{H}_2\text{Se}$  solution.

11. Write the equation in complete and correct the properties of acids & bases.

Acids	Bases
<input type="checkbox"/> Arrhenius	<input type="checkbox"/> Arrhenius
<input type="checkbox"/> Brønsted-Lowry	<input type="checkbox"/> Brønsted-Lowry
<input type="checkbox"/> Lewis	<input type="checkbox"/> Lewis
<input type="checkbox"/> Oxyacids	<input type="checkbox"/> Oxyacids
<input type="checkbox"/> Weak acids	<input type="checkbox"/> Weak bases

12. Classify each of the following acids & bases.

Name of Acid	Name of Base	Acid	Base
			Hydrochloric acid
			Barium hydroxide
		$\text{H}_2\text{SO}_4$	
		$\text{NaOH}$	
		$\text{H}_2\text{O}_2$	
			Hydrobromic acid

13. Predict the products for the following basic neutralization reaction involving each of these. Identify which compound is the best base (complete acid base) and complete base (partial acid base).

$$\text{NaOH} + \text{HCl} \rightarrow \boxed{\quad} + \boxed{\quad}$$
$$\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \boxed{\quad} + \boxed{\quad}$$
$$\text{LiOH} + \text{HCl} \rightarrow \boxed{\quad} + \boxed{\quad}$$

14. Label the following basic properties: acidic, basic, neutral, Bronsted Acid, Bronsted Base, Lewis Acid, Lewis Base.

15. pH scale of 0-14

16. Example:

# Chemistry Study Guide Acid And Bases Answers

**David R. Klein, Laurie S. Starkey**

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