



# CHEMISTRY

## CHEAT SHEET

# EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol  $\rightleftharpoons$ .

<p><b>What is Equilibrium?</b></p> <p>Chemical Equilibrium is a state where the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol <math>\rightleftharpoons</math>.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant <math>K_c</math> is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Equilibrium Constant (<math>K_p</math>)</b></p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant <math>K_p</math> is given by:</p> $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	
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<p><b>Le Chatelier's Principle: the effect of changing conditions on equilibrium</b></p>		
<p><b>Concentration</b></p>	<ul style="list-style-type: none"><li>Increasing concentration of a reactant or product shifts the equilibrium position to the right.</li><li>Decreasing concentration of a reactant or product shifts the equilibrium position to the left.</li></ul>	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the concentration of <math>N_2</math> or <math>H_2</math> shifts the equilibrium position to the right, increasing the yield of <math>NH_3</math>.</p>
<p><b>Pressure</b></p>	<ul style="list-style-type: none"><li>Increasing pressure shifts the equilibrium position to the side with fewer moles of gas.</li><li>Decreasing pressure shifts the equilibrium position to the side with more moles of gas.</li></ul>	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the pressure shifts the equilibrium position to the right, increasing the yield of <math>NH_3</math>.</p>
<p><b>Temperature</b></p>	<ul style="list-style-type: none"><li>Increasing temperature shifts the equilibrium position to the endothermic direction.</li><li>Decreasing temperature shifts the equilibrium position to the exothermic direction.</li></ul>	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the temperature shifts the equilibrium position to the left, decreasing the yield of <math>NH_3</math>.</p>

# Chemistry Study Guide Chemical Equilibrium

**M Carnoy**



## **Chemistry Study Guide Chemical Equilibrium:**

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Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a

complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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**General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing,1999 Resources in Education ,1984 **Study Guide for E. Russell Hardwick's Introduction to Chemistry** Joan Bouillon,1984 **Introduction to Chemistry, Study Guide** T. R. Dickson,1999-07-07 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and chemical vision and refining problem solving skills Many chapters in this edition has been rewritten and rearranged to vitalize the topics and to include interesting examples analogies and images

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**Differentiated Study Guide High School Chemistry** Carl David Morneweck,Chester Arthur Buckner,Gale R. Kirschner,1935 **Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

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Chemistry John Dickson, 1979-01-25      **Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed**  
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CHEMICAL PRINCIPLES ATOMS MOLECULES AND MOLES THE GAS LAWS AND THE ATOMIC THEORY MATTER WITH A  
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COVALENCE      **Chemical Principles** Steven S. Zumdahl, 1995 The Study Guide reflects the unique problem solving  
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