

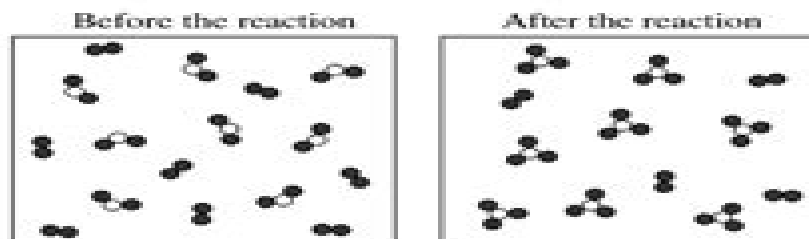
**CHEM 108**  
**ChemActivity 17: Limiting Reagents**  
**Solution Key**

**The Model: Limiting Reagent**

Sulfur dioxide reacts with oxygen to yield sulfur trioxide according to the equation:



While what's on the right may not be realistic, suppose there's a vessel that initially contains 8  $\text{SO}_2$  molecules and 8  $\text{O}_2$  molecules. The vessel is closed so that molecules cannot escape. If the reaction proceeds to the fullest extent, the "before" and "after" freeze frames of the system look like what's on the right.



**Key Questions**

- 1a. Refer to the Model. Which reactant –  $\text{SO}_2$  or  $\text{O}_2$  – remains in the "after" picture (and is, therefore, the **excess reagent**)? (Circle your answer.)
- b. Had there been more of a particular reactant, more sulfur trioxide would have been produced. Which reactant –  $\text{SO}_2$  or  $\text{O}_2$  – limited the production of sulfur trioxide? (Circle your answer.)
- c. In your own words, define what a **limiting reagent** is.

The limiting reagent is the reactant that limits the production of the product of a chemical reaction. If more limiting reagent is available, then more product can be generated by its consumption.

# Chemactivity 6 Alien

**Robert C. Smoot, Smoot, Richard G.  
Smith, Jack Price**



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*Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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