

Answer Key for Chemical Equilibria

1. Equilibrium constant expressions are always written as showing a ratio of products over reactants. Here is a general format for a generic equilibrium:



The equilibrium constant, K_c , has the ratio expressed as molar concentrations:

$$K_c = \frac{[A]^a [B]^b}{[C]^c [D]^d}$$

The equilibrium constant, K_c , is called an analytical equilibrium constant and it is applied to an analytical equilibrium constant expression. The analytical equilibrium constant expression will only include those components in the equilibrium that are in the same phase. Solids are never included because one cannot measure the concentration of a solid.

The equilibrium constant, K_p , is applied to gas phase equilibria. The equilibrium constant expression is expressed as a ratio in terms of the partial pressures of the gaseous components. Applying this to the same equilibrium above, assuming A, B, C and D are all gases:

$$K_p = \frac{P_C^c \cdot P_D^d}{P_A^a \cdot P_B^b}$$

Use these concepts to write the equilibrium constant expressions for the equilibria in this question:



This is a homogeneous equilibrium, meaning all the components are in the same phase, in this case, the gas phase. Thus, we can write both a K_c expression and a K_p expression:

$$K_c = \frac{[N_2H_4]}{[N_2] [H_2]^2} \quad \text{and} \quad K_p = \frac{P_{N_2H_4}}{P_{N_2} \cdot P_{H_2}^2}$$



This is also a homogeneous equilibrium, thus we can write both a K_c expression and K_p expression:

$$K_c = \frac{[PF_5] [F_2]}{[PF_3]} \quad \text{and} \quad K_p = \frac{P_{PF_5} \cdot P_{F_2}}{P_{PF_3}}$$



This is a heterogeneous equilibrium. Solids are never included in the equilibrium constant expressions, so we will only include the gases in this equilibrium. Since gases are involved, we are able to write both a K_c and K_p equilibrium constant expression:

$$K_c = \frac{[Cl_2]^2}{[SiCl_4] [H_2O]^2} \quad \text{and} \quad K_p = \frac{P_{Cl_2}^2}{P_{SiCl_4} \cdot P_{H_2O}^2}$$

Chemical Equilibrium Expression Answer Key

Daniel C. Harris



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