

For a first order reaction, $t_{1/2} = \frac{0.6932}{k}$

$$k = \frac{0.6932}{6.932 \times 10^4}$$

$$k = 10^{-5} \text{ s}^{-1}$$

$$k = \left(\frac{1}{t} \right) \ln \left(\frac{[A_0]}{[A]} \right)$$

$$10^{-5} \text{ s}^{-1} \times 100 \times 60 \text{ s} = \ln \left(\frac{[A_0]}{[A]} \right)$$

$$0.06 = \ln \left(\frac{[A_0]}{[A]} \right)$$

$$\frac{[A_0]}{[A]} = e^{0.06}$$

$$\frac{[A_0]}{[A]} = 1.06$$

$$\therefore \frac{[A_0] - [A]}{[A_0]} \times 100 \%$$

$$= \left(1 - \frac{[A]}{[A_0]} \right) \times 100 \%$$

$$= \left(1 - \frac{1}{1.06} \right) \times 100 \%$$

$$= 5.6 \%$$

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Table of Contents Chemical Kinetics Problems And Solutions

1. Understanding the eBook Chemical Kinetics Problems And Solutions
 - The Rise of Digital Reading Chemical Kinetics Problems And Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Kinetics Problems And Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Kinetics Problems And Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Kinetics Problems And Solutions
 - Personalized Recommendations
 - Chemical Kinetics Problems And Solutions User Reviews and Ratings
 - Chemical Kinetics Problems And Solutions and Bestseller Lists
5. Accessing Chemical Kinetics Problems And Solutions Free and Paid eBooks
 - Chemical Kinetics Problems And Solutions Public Domain eBooks
 - Chemical Kinetics Problems And Solutions eBook Subscription Services
 - Chemical Kinetics Problems And Solutions Budget-Friendly Options
6. Navigating Chemical Kinetics Problems And Solutions eBook Formats

- ePub, PDF, MOBI, and More
- Chemical Kinetics Problems And Solutions Compatibility with Devices
- Chemical Kinetics Problems And Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Kinetics Problems And Solutions
 - Highlighting and Note-Taking Chemical Kinetics Problems And Solutions
 - Interactive Elements Chemical Kinetics Problems And Solutions
- 8. Staying Engaged with Chemical Kinetics Problems And Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Kinetics Problems And Solutions
- 9. Balancing eBooks and Physical Books Chemical Kinetics Problems And Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Kinetics Problems And Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Kinetics Problems And Solutions
 - Setting Reading Goals Chemical Kinetics Problems And Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Kinetics Problems And Solutions
 - Fact-Checking eBook Content of Chemical Kinetics Problems And Solutions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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