

# Chemistry 110 Study Guide Chapters 55-57

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## Chapter 55

### *Key Concepts:*

- 1. The electronic structure of atoms determines their chemical properties.*
- 2. The types of elements are determined by the number of protons in the nucleus.*
- 3. The number of protons in an element's nucleus determines how strong the atom's nucleus is held together (termed "atomic number").*
- 4. The types of elements are determined by the number of electrons in the outermost energy level of the atom (termed "atomic number").*
- 5. Electrons in atoms are arranged in shells around the nucleus.*
- 6. The shells are filled in order of increasing energy.*
- 7. The first shell can hold a maximum of 2 electrons, the second shell can hold a maximum of 8 electrons, and so on.*
- 8. The outermost shell is called the valence shell.*
- 9. The valence shell determines how an atom will interact with other atoms.*
- 10. Atoms interact with each other in order to achieve a full valence shell.*
- 11. The types of elements are determined by the number of valence electrons in the outermost shell of the atom.*
- 12. The number of valence electrons determines how an atom will interact with other atoms.*
- 13. Atoms with the same number of valence electrons will interact with each other in the same way.*

### *Important Events:*

- 1. The discovery of the electron in 1897 by J. J. Thomson.*
- 2. The discovery of the nucleus in 1911 by Ernest Rutherford.*
- 3. The discovery of the proton in 1918 by Ernest Rutherford.*
- 4. The discovery of the neutron in 1932 by James Chadwick.*
- 5. The development of the periodic table of the elements by Dmitri Mendeleev in 1869.*

### *Notable Figures:*

- 1. J. J. Thomson - discovered the electron in 1897.*
- 2. Ernest Rutherford - discovered the nucleus in 1911 and the proton in 1918.*

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**AN Whitehead**



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1. Understanding the eBook Chemistry 110 Study Guide
  - The Rise of Digital Reading Chemistry 110 Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 110 Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry 110 Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 110 Study Guide
  - Personalized Recommendations
  - Chemistry 110 Study Guide User Reviews and Ratings
  - Chemistry 110 Study Guide and Bestseller Lists
5. Accessing Chemistry 110 Study Guide Free and Paid eBooks
  - Chemistry 110 Study Guide Public Domain eBooks
  - Chemistry 110 Study Guide eBook Subscription Services
  - Chemistry 110 Study Guide Budget-Friendly Options

6. Navigating Chemistry 110 Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry 110 Study Guide Compatibility with Devices
  - Chemistry 110 Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry 110 Study Guide
  - Highlighting and Note-Taking Chemistry 110 Study Guide
  - Interactive Elements Chemistry 110 Study Guide
8. Staying Engaged with Chemistry 110 Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry 110 Study Guide
9. Balancing eBooks and Physical Books Chemistry 110 Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry 110 Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 110 Study Guide
  - Setting Reading Goals Chemistry 110 Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 110 Study Guide
  - Fact-Checking eBook Content of Chemistry 110 Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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