

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release H^+ ions (protons)

Bases dissociate to accept H^+ ions (protons)

$$\text{Acid: } \text{H}_3\text{O}^+ \text{ (one } \text{H}^+ \text{ ion)} \quad \text{Base: } \text{OH}^- \text{ (one } \text{H}^+ \text{ ion)}$$

Acids dissociate to release H^+ ions, $\text{H}^+ + \text{base} \rightarrow \text{salt}$

What does H^+ ions, acids and bases do?

Acids and Bases

Acids are proton donors. $\left\{ \begin{array}{l} \text{The substance is losing its } \text{H}^+ \\ \text{The substance is giving up } \text{H}^+ \end{array} \right.$



What causes dissociation of acid-base reactions?

1) At the same time, the base participating change of the base and reacts to form acidic hydrogen of the acid.

2) At the end of the hydrogen and form the electrons in the other area.

Remember: $\text{acid} + \text{base} \rightarrow \text{salt} + \text{water}$

3) Both dissociation and dissociation are from Proton Properties. What if there is no proton available?

Acids, Bases, and Salts

Acids react to release H^+ ions. Bases react to release OH^- ions.



4) In the acid-base reaction, acids and bases are described in terms of strength.

5) Some of the acid-base reactions, according to my mother, are a potential partner, reaction with H_2 and form H_2O .

Acid Strength

What does a strong "strong acid"?

Generally speaking, the more the acid dissociates from water, the stronger it is.



For dissociation of the acid, a dissociation constant is described by dissociation constant.

$$\text{K}_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{H}_3\text{O}^+]} \quad \text{The stronger the acid, the larger the } \text{K}_a \text{ value.}$$

Smaller K_a value \rightarrow Stronger acid.



Acid Strength and pH

The pH value of strong acids (1) and strong bases (2) are shown below:

$$\text{pH}_1 = \frac{(-\log[\text{H}_3\text{O}^+])}{(-\log[\text{H}_3\text{O}^+])} = 10^{-1} = \text{pH}_2 = \frac{(-\log[\text{OH}^-])}{(-\log[\text{OH}^-])} = 10^{-1}$$

It is stated in a stronger acid because the pH is higher.

To simplify the notation, negative log value is the pH of the given acid.

$$\text{pH} = -\log[\text{H}_3\text{O}^+] = -\log[\text{H}_3\text{O}^+] = \text{pH}$$

As the stronger the acid, the lower the pH value. Or to express the fact:

Acids are pH values of more common chemical groups:



What makes the Acid Strong?

The strength of the acid is determined by the number of the conjugate base for the same when the acid dissociates. **The stronger the acid, the more acidic its conjugate base.**



Water is a weak acid and dissociates into protonated water and

Water is a very strong acid and dissociates into protonated water and

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Water is more dissociative than H_3O^+ in the same character, the stronger the acid

Therefore, the **water** dissociates \rightarrow **stronger** \rightarrow **stronger**

The protonated water dissociates \rightarrow **stronger** \rightarrow **stronger**

Acids \rightarrow also dissociates its conjugate base



The Position of Equilibrium for an Acid-Base Reaction

1) At equilibrium, no change occurs.



2) Acid dissociation equilibrium is another way of solving the problem:

In the expression of the reaction position, the right products is to the left reactants. To solve the position you need to consider that the equilibrium of the acid dissociation is shifted towards the weaker acid and the base.

The equilibrium position of the acid dissociation is to the right. Therefore, the equilibrium will be shifted to the higher pH (stronger acid).

3) When comparing for two acids, check their conjugate base: **Conjugate base**



4) If the pH of the conjugate acid is greater than the pH of the conjugate base, the equilibrium is shifted to the right. Therefore, the right position is to the stronger acid.



Choosing a Base to Dependent a Given Compound

Suggest you need to Dependent the following compound \rightarrow NaOH

Dependent bases common to weak acidic proton

Weak Acid conjugate base \rightarrow $\text{NaOH} \rightarrow \text{H}_2\text{O}$

From the chemical equation of the compound reacting with conjugate base of NaOH

$$\text{NaOH} + \text{H}_2\text{O} \rightleftharpoons \text{NaOH} + \text{H}_2\text{O} \quad \text{pH} = 14 \quad \text{pH} = 14$$

The pH of the conjugate acid must be higher than 14, so weaker acid is found.

Weak compound from a salt with a pH < 14. Another are weak acids (pH < 14).

$$\text{NaOH} + \text{H}_2\text{O} \rightleftharpoons \text{NaOH} + \text{H}_2\text{O} \quad \text{pH} = 14 \quad \text{pH} = 14$$

Choosing an Acid to Protonate a Given Compound

Suggest you need conjugate to protonate aqueous acid \rightarrow H_2SO_4 (strong acid)

Acid is a very common base and needs to be protonated before being dependent

Acids that have the chemical equation of salt \rightarrow H_2O

$$\text{NaOH} + \text{H}_2\text{O} \rightleftharpoons \text{NaOH} + \text{H}_2\text{O} \quad \text{pH} = 14 \quad \text{pH} = 14$$

The pH of strong like conjugate acid of acid is 14. Therefore, the pH of H_2O must be weaker than 14 (stronger acids, weaker acids).

Acids with a pH < 14 are protonated acids (good for protonating to other because the pH is 14 or it is the stronger acid in acid-base balance).



Acid-base dissociation Equilibrium

The equation shows whether the compound with a given pH is in its acidic or basic equilibrium than in an aqueous solution depending on the pH of the solution.

$$\text{pH} = \text{pH}_1 + \log \frac{[\text{H}_3\text{O}^+]}{[\text{H}_2\text{O}]}$$

Any given and application of the acid-base dissociation equation

$$\begin{array}{l} \text{If the pH} < \text{pH}_1: \quad \text{If the pH} > \text{pH}_1: \quad \text{If the pH} = \text{pH}_1: \\ \rightarrow \log \frac{[\text{H}_3\text{O}^+]}{[\text{H}_2\text{O}]} < 0 \quad \rightarrow \log \frac{[\text{H}_3\text{O}^+]}{[\text{H}_2\text{O}]} > 0 \quad \rightarrow \log \frac{[\text{H}_3\text{O}^+]}{[\text{H}_2\text{O}]} = 0 \end{array}$$

$$\begin{array}{l} \text{If } [\text{H}_3\text{O}^+] < [\text{H}_2\text{O}]: \quad \text{If } [\text{H}_3\text{O}^+] > [\text{H}_2\text{O}]: \quad \text{If } [\text{H}_3\text{O}^+] = [\text{H}_2\text{O}]: \\ \rightarrow \text{acidic state} \quad \rightarrow \text{basic state} \quad \rightarrow \text{neutral state} \end{array}$$

For example, imagine dissociation acidic salt to Dependent base due to its hydrolysis reaction.



The pH of conjugate hydrolysis acid is 1. Therefore, it is acidic. If the acidic dissociation reaction is stronger (pH < 1), acidic or the concentration of the acid and pH < 1, the drug will be in its acidic form.

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Michael B. Smith

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Columbia Review High-yield General Chemistry Stephen D. Bresnick, 1996 A study guide in question and answer format for basic chemistry. **Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry** Daniel C. Pantaleo, 1986 *Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney* William H. Myers, William T. Mooney, 1980

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