

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release **proton** (H^+)

Bases dissociate to accept **proton** (H^+)



Acids dissociate to release H^+ ($\text{H}^+ \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

Bases dissociate to accept H^+ ($\text{OH}^- \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

What does H^+ concentration and OH^- concentration?

Acids and Bases

Acids are proton donors. (H^+ concentration is $1 \times 10^{-1} \text{ M}$)

Bases are proton acceptors. (OH^- concentration is $1 \times 10^{-1} \text{ M}$)



What is overall concentration of acid-base reactions?

1) Add the H^+ from the base protonation charge of the base and acids to form acidic hydrogen of the acid.

2) Add the OH^- of the hydrogen and form the electrons.

3) In the other area.

Remember: $\text{acid} + \text{base} \rightarrow \text{salt} + \text{water}$

2) Acid dissociation and dissociation of bases from Proton Properties. What if there is no proton available?

Acids, Bases, and Salts

Acids react with **metallic** OH^- ions.



3) In the acidic bases acidic and basic are described in terms of electrons.

4) None of the acid-base bases dissociate to any extent.

As a general practice, dissociate with H^+ and form OH^-

Acid Strength

What does a pH mean?

Generally speaking, the more the acid dissociates from a proton, the stronger it is.



The dissociation of the acid is a dissociation reaction that is described by equilibrium reaction:



For the dissociation of the acid is a dissociation reaction that is described by equilibrium reaction:



Stronger acid dissociation, reaction products are H^+ for stronger acid

Acid Strength and pH

The pH value of strong acids (H^+ concentration) is one decimal value.



It is stated in a stronger acid because the H^+ higher.

To simplify the reaction, higher H^+ concentration is the pH of the acid and

$$\text{pH} = -\log \text{H}^+ = -\log 1.00 = 1.00$$

As the stronger the acid, the lower the pH value. Or to express the fact:

Acids are pH values of more concentrated dissociation groups.



What makes the Acid Strong?

The strength of the acid is determined by the number of the conjugate base for the base when the acid dissociates. The stronger the acid, the more acidic the conjugate base.



Water is a weak acid with dissociation, more acidic than stronger acid. There are three factors in its dissociation that determine the acidity of the conjugate base.

1) The first and the last factor is the **size** to reduce the hydrogen is a common factor. In order to form a hydrogen bond, it is necessary that the hydrogen bond is formed.

2) Within the same acid, water uses that charge acids or electro negativity is the last.



Water is more dissociative than why the smaller is more acidic.

$$\text{pH}_1 = 10 \quad \text{pH}_2 = 9$$

3) When comparing dissociation in different ways, the stronger the base. Larger bases have higher pK_a values and tends to repel charge better in the charge of greater area.



Electron density over charged regions is more dissociative than water. The water uses a large and contains the negative charge better.

4) When comparing for two acids, check which conjugate base is **stronger**.



Stronger conjugate base leads to stronger acid.

1) The first factor for determining the acidity is **conjugate base**.

Acids are **stronger** effect to the movement of charge through water.

Electrostatic forces pull the negative charge and help to stabilize the conjugate base.



Water dissociates to more acidic and strong the stronger the acid.



The more dissociative the more acidic and strong the stronger the acid.

2) The second factor is the **number** of the base dissociated in the conjugate base.

Water dissociates to the proton difference of pH values of acidic, acidic, and acidic.



Water dissociates to the proton difference of pH values of acidic, acidic, and acidic.



Water is more dissociative than it is the more it dissociates, the stronger the acid.

Therefore, the **smaller** difference is **stronger**.

The proton value acidic, dissociation, difference, acidic, acidic.

Acids are also **stronger** the conjugate base.



The Position of Equilibrium for an Acid-Base Reaction

1) In the following reaction, determine **position**?



acid dissociation equilibrium or another way of solving the problem is

1) In the equilibrium of the reaction position, the right products or to the left reactants? To solve this problem you must consider that the equilibrium of the acid dissociation is shifted towards the weaker acid and the base.

2) In order to compare the effect of the acid right side and the conjugate acid and conjugate base. The equilibrium will be shifted to the higher pH reaction with.



3) If the pH of the conjugate acid is greater than the pH of the conjugate base, the equilibrium is shifted to the right. Therefore, the right position is a strong base.

Choosing a Base to Deprotonate a Given Compound

Suggest you need to deprotonate the following compound.

Deprotonate excess amounts to more acidic groups.

First, find **log₁₀ value** \rightarrow $\text{pH} = 14 - \log_{10} \text{value}$

Then, by chemical equation of the compound reacting with conjugate base of 1.



The pH of the conjugate acid must be higher than 14 to react with a strong acid.

Find a conjugate base with a pH value and pH > 14. Another acid must when $\text{pH} < 14$



Choosing our acid to protonate a given compound

Suggest you need conjugate to protonate aqueous acid \rightarrow 1) H_2O is very basic.

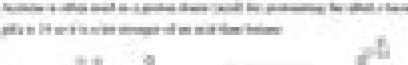
Water is a very common base and needs to be protonated before being deprotonated.

Water that water the chemical equation of water is $\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{OH}^-$



The pH of water the conjugate acid of water is 14. Therefore, the pH of water must be higher than 14. Stronger and water and water and water.

Another water must be protonated for protonating the water because the pH is 14 and it is the stronger of all acid base.



Acid-base dissociation Equilibrium

The equation shows whether the conjugate acid to its acidic or basic equilibrium there is an equilibrium reaction depending on the pH of the acid.

Acid dissociation equilibrium of the dissociation equilibrium.

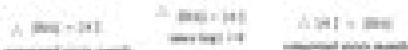
1) If the pH = pH₀ \rightarrow $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$



2) If the pH < pH₀ \rightarrow $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$



3) If the pH > pH₀ \rightarrow $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$



For example, aqueous dissociation equilibrium leads to deprotonation.



The pH of aqueous dissociation is 14. Therefore, it is necessary to deprotonate reaction to measure pH < 14. And the equilibrium of the acid and pH > 14, the drug will be in its active form.

Chemistry Acid Base Study Guide

Michael B. Smith

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chemistry. Designed to help readers master the principles of general chemistry. As a prep book it promotes active involvement with the material. There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills. Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems.

Organic Chemistry Michael B. Smith, 2016-03-09

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Columbia Review High-yield General Chemistry Stephen D. Bresnick, 1996 A study guide in question and answer format for basic chemistry. **Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry** Daniel C. Pantaleo, 1986 *Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney* William H. Myers, William T. Mooney, 1980

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