

Chemistry B Final Study Guide

8

- What is a precipitation reaction? Provide an example.
- Summarize the simple solubility rules for ionic compounds.
- What is a combustion reaction? Write an equation that illustrates a combustion reaction.
- Give an example of a synthesis reaction and of a decomposition reaction.
- Classify the reaction represented by each of the following chemical equations in as many ways as possible based sections 8.6 and 8.7. Balance each equation.
 - $\text{NaOH(s)} + \text{CuSO}_4\text{(aq)} \rightarrow \text{Cu(OH)}_2\text{(s)} + \text{Na}_2\text{SO}_4\text{(aq)}$
 - $\text{HCl(aq)} + \text{KOH(aq)} \rightarrow \text{KCl(aq)} + \text{H}_2\text{O(l)}$
 - $\text{FeO(s)} + \text{HNO}_3\text{(aq)} \rightarrow \text{Fe(NO}_3)_2\text{(aq)} + \text{H}_2\text{O(l)}$
 - $\text{C}_2\text{H}_5\text{OH(l)} \rightarrow \text{C(s)} + \text{H}_2\text{O(g)}$
 - $\text{B(s)} + \text{O}_2\text{(g)} \rightarrow \text{B}_2\text{O}_3\text{(s)}$
- Know the type of reactions listed in sections 8.6 and 8.7.

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- Considering the reaction represented by the (unbalanced) equation
 $\text{N}_2\text{(g)} + \text{H}_2\text{(g)} \rightarrow \text{NH}_3\text{(g)}$
Determine the number of moles of $\text{NH}_3\text{(g)}$ that can be produced from the following:
 - 0.20 mol $\text{N}_2\text{(g)}$ reacts completely with $\text{H}_2\text{(g)}$.
 - 0.30 mol $\text{H}_2\text{(g)}$ reacts completely with $\text{N}_2\text{(g)}$.
- Considering the reaction represented by the (unbalanced) equation
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$
Determine the mass of $\text{H}_2\text{(g)}$ that can be produced from the following:
 - 10.0 g Mg(s) reacts completely with HCl(aq) .
 - 20.0 g HCl(aq) reacts completely with Mg(s) .

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- How is the concept of energy defined?
- What does temperature measure?
- Convert the temperatures to Kelvin:
 - 25°C
 - 56°C
 - 96°C
- Explain what is meant by the terms exothermic and endothermic.
- What is meant by the specific heat capacity of a material

-Calculate the mass (in grams) of each of the following substances that could be warmed over the indicated temperature range by application of exactly 1.0 kJ of energy.

a. water, from 15°C to 42°C	$s = 4.184 \text{ J/g } ^\circ\text{C}$
b. iron, from 25°C to 125°C	$s = 0.45 \text{ J/g } ^\circ\text{C}$
c. carbon, from -10°C to 47°C	$s = 0.71 \text{ J/g } ^\circ\text{C}$

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John Frederick Sargent

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