

CHAPTER

02

SOLUTION

- Solution :** A homogeneous mixture of two or more substances is known as solution.
- Solute :** The substance present in smaller amount in a solution is called solute.
- Solvent :** The substance present in larger amount in a solution is called solvent.
- Binary solution :** Those solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
- Ternary solutions :** Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
- Aqueous solution :** When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
- Non-aqueous solution :** When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
- Weight Fraction** = $\frac{\text{weight of solute}}{\text{weight of solution}}$
- Weight percentage**
= $\frac{\text{weight of solute}}{\text{weight of solution}} \times 100$
Weight of solute present in 100 grams of solution is called weight percentage
- Volume of solute present in 100 ml of the solution is called volume percentage

- Volume percentage**
= $\frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$
- Volume Fraction** = $\frac{\text{Volume of solute}}{\text{Volume of solution}}$
- Molarity (M)**
= $\frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$
(or) $M = \frac{W}{GMW} \times \frac{1}{V(\text{lit})}$
- Molarity** = $\frac{\text{No of moles of solute}}{1000} \times \frac{1000}{\text{Volume of solution in ml}}$
- Molarity** = $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{\text{V in ml}}$
(or) $M = \frac{W}{GMW} \times \frac{1000}{V(\text{ml})}$
- Weight of solute** = $\frac{M \times GMW \times V(\text{ml})}{1000}$
- Molarity for dilute solutions** : $M_1V_1 = M_2V_2$
- For neutralisation reactions** : $\frac{M_1V_1}{n_1} = \frac{M_2V_2}{n_2}$
- For mixed solution molarity**
= $\frac{M_1V_1 + M_2V_2 + \dots}{V_1 + V_2 + \dots}$

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