

## CHAPTER

## 02

## SOLUTION

1. **Solution** : A homogeneous mixture of two or more substances is known as solution.
2. **Solute** : The substance present in smaller amount in a solution is called solute.
3. **Solvent** : The substance present in larger amount in a solution is called solvent.
4. **Binary solution** : Those solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
5. **Ternary solutions** : Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
6. **Aqueous solution** : When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
7. **Non-aqueous solution** : When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
8. **Weight Fraction** =  $\frac{\text{weight of solute}}{\text{weight of solution}}$
9. **Weight percentage**  

$$= \frac{\text{weight of solute}}{\text{weight of solution}} \times 100$$

Weight of solute present in 100 grams of solution is called weight percentage

Volume of solute present in 100 ml of the solution is called volume percentage
10. **Volume percentage**  

$$= \frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$$
11. **Volume Fraction** =  $\frac{\text{Volume of solute}}{\text{Volume of solution}}$
12. **Molarity (M)**  

$$= \frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$$

$$(\text{or}) M = \frac{W}{GMW} \times \frac{1}{V(\text{lit})}$$
13. **Molarity** =  $\frac{\text{No of moles of solute}}{\text{Volume of solution in ml} \times \frac{1000}{1000}}$
14. **Molarity** =  $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{V \text{ in ml}}$   

$$(\text{or}) M = \frac{W}{GMW} \times \frac{1000}{V(\text{ml})}$$
15. **Weight of solute** =  $\frac{M \times GMW \times V \text{ in ml}}{1000}$
16. **Molarity for dilute solutions** :  $M_1V_1 = M_2V_2$
17. **For neutralisation reactions** :  $\frac{M_1V_1}{n_1} = \frac{M_2V_2}{n_2}$
18. **For mixed solution molarity**  

$$= \frac{M_1V_1 + M_2V_2 + \dots}{V_1 + V_2 + \dots}$$

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## **Table of Contents Chemistry Dimensions 2 Solutions**

1. Understanding the eBook Chemistry Dimensions 2 Solutions
  - The Rise of Digital Reading Chemistry Dimensions 2 Solutions
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Dimensions 2 Solutions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Dimensions 2 Solutions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Dimensions 2 Solutions

- Personalized Recommendations
- Chemistry Dimensions 2 Solutions User Reviews and Ratings
- Chemistry Dimensions 2 Solutions and Bestseller Lists
- 5. Accessing Chemistry Dimensions 2 Solutions Free and Paid eBooks
  - Chemistry Dimensions 2 Solutions Public Domain eBooks
  - Chemistry Dimensions 2 Solutions eBook Subscription Services
  - Chemistry Dimensions 2 Solutions Budget-Friendly Options
- 6. Navigating Chemistry Dimensions 2 Solutions eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Dimensions 2 Solutions Compatibility with Devices
  - Chemistry Dimensions 2 Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Dimensions 2 Solutions
  - Highlighting and Note-Taking Chemistry Dimensions 2 Solutions
  - Interactive Elements Chemistry Dimensions 2 Solutions
- 8. Staying Engaged with Chemistry Dimensions 2 Solutions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Dimensions 2 Solutions
- 9. Balancing eBooks and Physical Books Chemistry Dimensions 2 Solutions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Dimensions 2 Solutions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Dimensions 2 Solutions
  - Setting Reading Goals Chemistry Dimensions 2 Solutions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Dimensions 2 Solutions

- Fact-Checking eBook Content of Chemistry Dimensions 2 Solutions
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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