

**Chemistry | Semester 2 | Final Exam Study Guide**  
**Unit 9: Compound Stoichiometry**

38. A mole is a unit defined to be  $= 6.02 \times 10^{23}$  particles  
 1 mole = mass

39. The molar mass of an element is the mass of one mole of the substance  
 = its atomic mass in grams

40. What information is given to you in the empirical formula?

- elements in comp.
- ratio of elements

41. What do the subscripts represent in a molecular formula?  
 how many atoms of each element

42. Define Formula Mass  
 mass of 1 mole of an ionic compound

43. Define Molar Mass  
 mass of 1 mole of any compound

44. Define Molecular Mass  
 mass of one mole of a covalent (molecular) comp.

True/False

45. F One mole of  $O_2$  has a greater mass than 1 mole of  $Na_2O$ .  
 32 vs 61.98
46. F One mole of  $H_2$  has <sup>same</sup> molecules than 1 mole of  $Na_2O$ .  
 same # molecules
47. F One molecule of  $H_2$  has more atoms than 1 molecule of  $Na_2O$ .  
 2 atoms vs 3 atoms

**Calculations**

48. Find the molar mass of  $Al_2(SO_4)_3$

Find mass of each element  $\rightarrow$  multiply each mass by subscript  $\rightarrow$  add masses of all elements together = molar mass

$$\begin{array}{l} Al \quad (2) \quad 26.98 = 53.96 \\ S \quad (3) \quad 32.07 = 96.21 \\ O \quad (12) \quad 16.00 = 192.00 \end{array} \quad 294.17 \text{ g/mol}$$

49. How many moles are in 2.4g of carbon dioxide ( $CO_2$ )?

Start with the given value  $\rightarrow$  use dimensional analysis so you can factor out labels until you get to target units

$$\frac{2.4 \text{ g } CO_2}{44 \text{ g } CO_2} \times \frac{1 \text{ mol } CO_2}{1 \text{ mol } CO_2} = 0.05 \text{ mol}$$

molar mass  $CO_2$

50. How many molecules are contained in 35.37g of water (molar mass of water is 18.02g/mol)?

$$\frac{35.37 \text{ g } H_2O}{18.02 \text{ g } H_2O} \times \frac{1 \text{ mol } H_2O}{1 \text{ mol } H_2O} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol } H_2O} = 1.18 \times 10^{24} \text{ molecules } H_2O$$

51. How many atoms are in 6.3 moles of calcium?

Hint: there are  $6.02 \times 10^{23}$  atoms in 1 mole

$$\frac{6.3 \text{ mol Ca}}{1 \text{ mol Ca}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol Ca}} = 3.79 \times 10^{24} \text{ atoms}$$

52. Calculate the number of moles there are in  $4.5 \times 10^{26}$  molecules of sulfur dioxide ( $SO_2$ ).

$$\frac{4.5 \times 10^{26} \text{ molecules } SO_2}{6.02 \times 10^{23} \text{ molecules}} \times \frac{1 \text{ mol } SO_2}{1 \text{ mol } SO_2} = 747.51 \text{ mol } SO_2$$

53. Calculate the percent composition of sodium and oxygen in sodium oxide ( $Na_2O$ ).

Find the molar mass  $\rightarrow$  divide total mass of element where it's solving for by the total mass of compound

$$\begin{array}{l} Na \quad (2) \quad 22.99 = 45.98 / 61.98 = 74.2\% \\ O \quad (1) \quad 16.00 = 16.00 / 61.98 = 25.8\% \end{array}$$

molar mass

$$\begin{array}{r} 45.98 \\ 16.00 \\ \hline 61.98 \text{ g/mol} \end{array}$$

54. What is the empirical formula of a compound that contains 81.68% carbon and 18.32% hydrogen?

Assume 100 g: % to mass  $\rightarrow$  mass to mole  $\rightarrow$  divide by smallest  $\rightarrow$  multiply till whole

$$\frac{81.68 \text{ g C}}{12.01 \text{ g C}} \times \frac{1 \text{ mol C}}{1 \text{ mol C}} = 6.8 / 6.8 = 1 \times 3 = 3$$

$$\frac{18.32 \text{ g H}}{1.01 \text{ g H}} \times \frac{1 \text{ mol H}}{1 \text{ mol H}} = 18.139 / 6.8 = 2.67 \times 3 = 8$$

**$C_3H_8$**

55. What is the molecular formula of a hydrocarbon that has an empirical formula of  $CH_4$  and a molecular mass (molar mass of molecule) of 48.15 g/mol?

$$\frac{MM_{\text{molecule}}}{MM_{\text{empirical}}} = \frac{48.15}{16.05} = 3 \quad \begin{array}{l} C_1 H_4 \\ \times 3 \\ \hline C_3 H_{12} \end{array}$$

Multiply each subscript by the above answer to get the subscripts for the molecular formula

To find Empirical Formula

divide (multiply till whole)

if it ends in	multiply by
0.2 or 0.8	$\times 5$
0.25 or 0.75	$\times 4$
0.33 or 0.67	$\times 3$
0.5	$\times 2$

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## **Table of Contents Chemistry Final Exam Study Guide Key**

1. Understanding the eBook Chemistry Final Exam Study Guide Key
  - The Rise of Digital Reading Chemistry Final Exam Study Guide Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Final Exam Study Guide Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Final Exam Study Guide Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Final Exam Study Guide Key
  - Personalized Recommendations
  - Chemistry Final Exam Study Guide Key User Reviews and Ratings
  - Chemistry Final Exam Study Guide Key and Bestseller Lists
5. Accessing Chemistry Final Exam Study Guide Key Free and Paid eBooks
  - Chemistry Final Exam Study Guide Key Public Domain eBooks
  - Chemistry Final Exam Study Guide Key eBook Subscription Services
  - Chemistry Final Exam Study Guide Key Budget-Friendly Options
6. Navigating Chemistry Final Exam Study Guide Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Final Exam Study Guide Key Compatibility with Devices
  - Chemistry Final Exam Study Guide Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Final Exam Study Guide Key
  - Highlighting and Note-Taking Chemistry Final Exam Study Guide Key
  - Interactive Elements Chemistry Final Exam Study Guide Key
8. Staying Engaged with Chemistry Final Exam Study Guide Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Final Exam Study Guide Key
- 9. Balancing eBooks and Physical Books Chemistry Final Exam Study Guide Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Final Exam Study Guide Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Final Exam Study Guide Key
  - Setting Reading Goals Chemistry Final Exam Study Guide Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Final Exam Study Guide Key
  - Fact-Checking eBook Content of Chemistry Final Exam Study Guide Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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