

2nd Nine Weeks

Unit 4 & 5: Chemical Bonds & VSEPR

1. Identify what type of bond has the following properties.

(Given bond: Metallic, Ionic, Covalent)

<u>Metallic</u> Bond	<u>Ionic</u> Bond	<u>Covalent</u> Bond
- Sea of delocalized electrons - Conducting when solid	- Crystalline solids - High melting & boiling point - Conducts electricity when dissolved in solution	- Can be solid, liquid, or gas - Lower melting & boiling point - Does not conduct electricity

2. Circle the correct word that will complete the sentences below.

Ionic bonds have a large (small) electronegativity difference. The element with a stronger electronegativity will give (take) electrons from the element with a weaker (stronger) electronegativity.

Covalent bonds have a large (small) electronegativity difference. So, the elements will give (take) their electrons with each other to fill (lose) the outer (inner) shell(s) rule.

3. Fill in the chart below using the correct bond or Covalent Rule.

Bond or Covalent?	Formula	Name	Lewis Structure	if Covalent Polarity	if Covalent Shape
<u>Ionic</u>	<u>Na₂S</u>	<u>Sodium Sulfide</u>	<u>2[Na]⁺ [S]²⁻</u>	Non polar	Tetrahedral
<u>Covalent</u>	<u>CH₄</u>	<u>Carbon tetrahydride</u>	<u>H-C(H)(H)(H)</u>	<u>Non polar</u>	<u>Tetrahedral</u>
<u>Ionic</u>	<u>CuNO₃</u>	<u>Copper (I) nitrate</u>	<u>[Cu]⁺ 3 [NO₃]⁻</u>	Non polar	Tetrahedral
<u>Covalent</u>	<u>OF₂</u>	<u>Oxygen difluoride</u>	<u>F-O-F</u>	<u>Polar</u>	<u>Bent</u>
<u>Ionic</u>	<u>NaOH</u>	<u>Sodium Hydroxide</u>	<u>[Na]⁺ [OH]⁻</u>	Non polar	Tetrahedral
<u>Ionic</u>	<u>CuCN</u>	<u>Copper (I) cyanide</u>	<u>[Cu]⁺ [CN]⁻</u>	Non polar	Tetrahedral
<u>Covalent</u>	<u>BCl₃</u>	<u>Boron trichloride</u>	<u>Cl-B-Cl</u>	<u>Non polar</u>	<u>Trigonal Planar</u>
<u>Covalent</u>	<u>H₂S</u>	<u>Hydrogen monosulfide</u>	<u>H-S-H</u>	<u>Polar</u>	<u>Bent</u>

Unit 6: Electrons

Important Vocabulary & their Definition:

- Wavelength: length of one wave from top of one wave to another
- Frequency: number of waves

1. Circle the correct word that will complete the sentences below.

A wave with a large wavelength will have a high (low) frequency and high (low) energy.

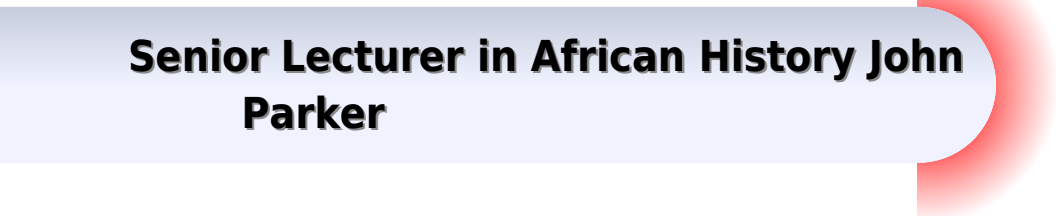
2. Calculate the energy (E) and wavelength (λ) of a photon of light with a frequency (f) of 6.465×10^{14} Hz.

$$\begin{aligned}
 E &= hf \\
 &= (6.63 \times 10^{-34}) (6.465 \times 10^{14}) \\
 E &= 4.28 \times 10^{-19} \text{ J} \\
 \lambda &= \frac{(6.63 \times 10^{-34}) (3 \times 10^8)}{4.28 \times 10^{-19}} = \boxed{4.66 \times 10^{-7} \text{ m}}
 \end{aligned}$$

$$\begin{aligned}
 h &= 6.63 \times 10^{-34} \\
 c &= 3 \times 10^8
 \end{aligned}$$

Chemistry Final Study Guide Answer Key

**Senior Lecturer in African History John
Parker**



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