

Worksheet: Neutralization: Chapter 21, (p. 612–630 only)**Page 1**

Complete and *balance* these neutralization reactions. Label the acid (starts with "H"), base (ends in "OH") and *name* the salt produced. Be sure to write the *correct formula* for the salt!

I've named the unfamiliar acids for you. Remember that -ic acids make -ate ions and only the FIRST H's are acidic in organic acids. (p. 613–615)



base

benzoic acid

Salt = magnesium benzoate

**Hydrolysis and Buffers** (p. 626–630)

9. Identify the acid, base, conjugate acid and conjugate base in these hydrolysis reactions.



10. Identify the following as Salts of strong or weak bases and strong or weak acids.
(HCl, HNO₃ and H₂SO₄ are strong acids. NH₃ [NH₄OH] is a weak base.)

Would their solutions in water be acidic, basic or neutral?

weak/strong base

weak/strong acid

acidic, basic or neutral solution?

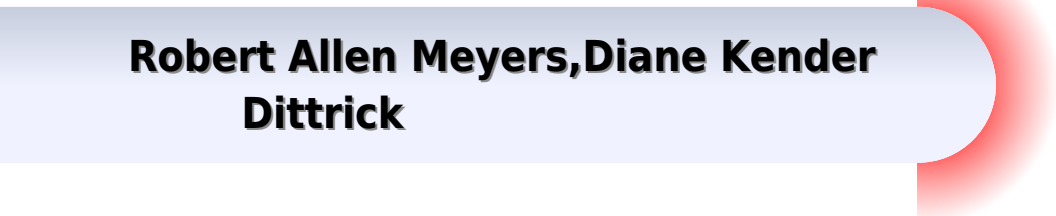


11. H₂CO₃ is a *much weaker* acid than HC₂H₃O₂. Would solutions of each of these *salts* be:

Neutral, a little Basic or Very Basic? NaCl _____ NaC₂H₃O₂ _____ Na₂CO₃ _____ Explain.

Chemistry Neutralization And Answer Key

**Robert Allen Meyers, Diane Kender
Dittrick**



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Meyers, Diane Kender Dittrick, 1999 This condensed version of the eight volume Encyclopedia of Environmental Analysis and Remediation provides a convenient source of information on environmental hazards and the methods that can be employed to detect and repair them General Chemistry Laboratory Text with Qualitative Analysis Wilbert Hutton, 1968 **Reprints from the Departments of Chemistry and Chemical Engineering of the University of Michigan** University of Michigan. Department of Chemistry, 1947

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Table of Contents Chemistry Neutralization And Answer Key

1. Understanding the eBook Chemistry Neutralization And Answer Key
 - The Rise of Digital Reading Chemistry Neutralization And Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Neutralization And Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Neutralization And Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Neutralization And Answer Key
 - Personalized Recommendations
 - Chemistry Neutralization And Answer Key User Reviews and Ratings
 - Chemistry Neutralization And Answer Key and Bestseller Lists

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 - Chemistry Neutralization And Answer Key Enhanced eBook Features
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 - Highlighting and Note-Taking Chemistry Neutralization And Answer Key
 - Interactive Elements Chemistry Neutralization And Answer Key
8. Staying Engaged with Chemistry Neutralization And Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Neutralization And Answer Key
9. Balancing eBooks and Physical Books Chemistry Neutralization And Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Neutralization And Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Neutralization And Answer Key
 - Setting Reading Goals Chemistry Neutralization And Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Neutralization And Answer Key
 - Fact-Checking eBook Content of Chemistry Neutralization And Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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