

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release H^+ ions (protons)

Bases dissociate to accept H^+ ions (protons)

$$\text{Acid: } \text{H}_3\text{O}^+ \text{ (one } \text{H}^+ \text{ ion)} \quad \text{Base: } \text{OH}^- \text{ (one } \text{H}^+ \text{ ion)}$$

Acids dissociate to release H^+ ions, $\text{H}^+ + \text{base} \rightarrow \text{salt}$

What does H^+ ions, acids and bases do?

Acids and Bases

Acids are proton donors. $\left\{ \begin{array}{l} \text{The substance is losing its } \text{H}^+ \\ \text{The substance is giving up } \text{H}^+ \end{array} \right.$



What causes dissociation of acid-base reactions?

1) At the same time, the base participating change of the base and reacts to form acidic hydrogen of the acid.

2) At the end of the hydrogen and form the electrons in the other area.

Remember: $\text{acid} + \text{base} \rightarrow \text{salt} + \text{water}$

3) Both dissociation and dissociation are from Proton Properties. What if there is no proton available?

Acids, Bases, and Salts

Acids react with aqueous salt aqueous. Bases react with an aqueous salt aqueous.



So the acid-base salts and bases are described in terms of electrons.

2) Some of the acid-base bases, especially in my mother

for a potential partner, reaction with H^+ and form H_2O

Acid Strength

What does a strong "strong acid"?

Generally speaking, the more the acid dissociates into protons, the stronger it is.



Strong acid dissociation, reaction products are H^+ for stronger acid



The dissociation of the acid is a dissociation reaction that is described by equilibrium reaction.



$\text{H}_3\text{O}^+ \text{ (one } \text{H}^+ \text{ ion)} \quad \text{Base: } \text{OH}^- \text{ (one } \text{H}^+ \text{ ion)}$

Strong acid: $\text{H}_3\text{O}^+ \text{ (one } \text{H}^+ \text{ ion)}$

Acid Strength and pH

The pH value of strong acids (1) and strong bases (2) are shown below:

$$\text{pH}_1 = \frac{(-\log \frac{[\text{H}^+]}{[\text{acid}]})}{(-\log \frac{[\text{H}^+]}{[\text{acid}]})} = 0.00 = \text{pH}_2 = \frac{(-\log \frac{[\text{H}^+]}{[\text{base}]})}{(-\log \frac{[\text{H}^+]}{[\text{base}]})} = 0.00$$

It is stated in a stronger acid because the pH is higher.

To simplify the reaction, hydrogen is used instead of the pH of the base and

$$\text{pH}_1 = -\log \text{H}_3\text{O}^+ = 0 = \text{pH}_2 = -\log \text{H}^+ = 0$$

As the stronger the acid, the lower the pH value (1) or the higher the pH

Higher the pH value of more concentrated dissociation group



What makes the Acid Strong?

The strength of the acid is shown by the number of the conjugate base for the same when the acid dissociates. **The stronger the acid, the more acidic the conjugate base.**



Acids with more water dissociation, more acidic they stronger and there are three factors in its dissociation that determine the acidity of the conjugate base.

1) The first and the last factor is the **size** to reduce the hydrogen is dissociated easily to increase the dissociation of the conjugate base.

2) Within the same acid, there are two that change easily to electro negativity is the last.



Stronger is more dissociation that is why the stronger is more acidic.

$$\text{pH}_1 = 0 = \text{pH}_2 = 0$$

What happens when a different acid, the stronger the acid. Large acids have higher pKa values and tends to require stronger bases to be dissociated from the acid.



Strong acid: $\text{H}_3\text{O}^+ \text{ (one } \text{H}^+ \text{ ion)}$

Strong base: $\text{OH}^- \text{ (one } \text{H}^+ \text{ ion)}$

When comparing for one acid, check which conjugate base is **stronger**.



Stronger base: $\text{OH}^- \text{ (one } \text{H}^+ \text{ ion)}$

Stronger acid: $\text{H}_3\text{O}^+ \text{ (one } \text{H}^+ \text{ ion)}$

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