

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent	solvent
insoluble	solid	solid	solid

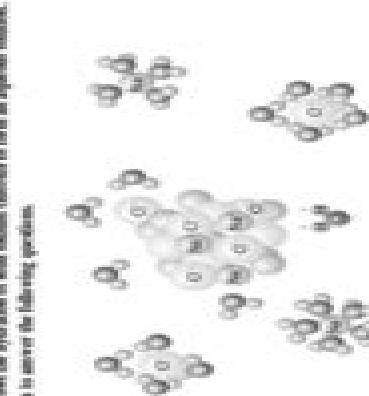
After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is the (3) _____ solvent _____ and solute is the

(4) _____ solution. A solution that does not dissolve is
 (5) _____ insoluble. (6) _____ liquid _____ solution is the most common
 type of solution. If one liquid is soluble in another liquid, such as water and oil, the
 two liquids are (7) _____ miscible. However, if one liquid is insoluble in another
 the liquids are (8) _____ immiscible.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.



The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Underline the diagram to answer the following questions.

1. In a solution, what is the solute in a salt? What is solute?
 2. In a solution, what is the process of surrounding solute particles with solvent particles to form a solution?

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Density, Matter and Change • Chapter 11

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CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Section 15.1** *Answers***STUDY GUIDE FOR CONTENT MASTERY**

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.
 17. Unlike the attraction of the water molecule toward the sodium ion and chloride ion,
 because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom), because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. Unlike the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.
 19. In a glass of water, what does the rate of solution may be increased?
 stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with their examples.

True 20. The overall energy change that occurs when a solute dissolves is called the heat of solution.

False 21. Solubility is a property of the solute, amount of either the solute or a given amount of solute at a specified temperature and pressure.

False 22. Solubility continues to change as the solute dissolves in the solvent because the solute particles are smaller.

True 23. In a saturated solution, solute and solvents are in equilibrium.

True 24. Additional solute can be dissolved in an unsaturated solution.

True 25. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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David Dimmlich

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web what does the crossed arrow represent in the comic strip 11 what are two of the definitions of dipole given in the comic strip 12 making sense 13 if you finish early what does electronegativity have to do with polarity using polar bears and penguins create an illustration showing a hydrogen sulfide molecule h₂s

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web each dipole represents one polar bond the molecule is nonpolar 10 describe what you think is happening to the penguins in the illustration of h₂o on page 4 they are being pulled towards the stronger polar bear he s more electronegative 11

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web the two polar bears or penguins are equal in electronegativity therefore they pull electrons equally as hard and neither one will end up with more electrons more of the time 4 a what 3 types of bonds are represented on the third page of the comic book nonpolar covalent polar covalent and ionic b

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web answer because the two polar bears or penguins have equal electronegativity they will pull electrons equally hard and in most cases neither will end up with more electrons each polar bear and penguin has an electronegativity of the same value a non polar bond is formed when two polar bears or two penguins are bonded together

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web why does the polar bear tell the penguin to try to be more positive the polar bear benefits more when the penguin is more positive the more the polar bear attracts the shared pair of electrons the more positive the partial charge on the penguin

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web answer the following questions consider the following illustration draw the lewis dot structure for hcl if the penguin represents a hydrogen atom and the polar bear represents a chlorine atom what does the ice cream represent in the drawing what do you think the picture is trying to illustrate would hcl be attracted to the charged wand

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web to answer this question you will examine 1 electronegativity 2 nonpolar vs polar molecules 3 electronegativity and bonding electronegativity the h atom and the cl atom in hydrogen chloride

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web 1 how does the comic strip define a polar molecule 2 define electronegativity as you understand it after reading the first two pages of the comic strip 3 what is the artist trying to represent by two polar bears arm wrestling or two penguins arm wrestling 4 what three types of bonds are represented on the third page of the comic strip

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web explain how the iceberg penguins and polar bears represent trends in electronegativity on the periodic table and more study with quizlet and memorize flashcards containing terms like how does the comic define a polar molecule define electronegativity as you understand it after reading the first two pages of the comic interpret the pic

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web so the final representation of the nh 3 molecule using polar bears and penguins is as follows begin array c text polar bear nitrogen text with two lone pairs of electrons text and three penguins hydrogen end array this is a fun and creative way to represent the nh 3 molecule

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web polar bears and penguins purpose in this lesson you will be exploring polarity and bonding between atoms in greater

detail a comic book will provide new information about these topics and will introduce you to the concept of electronegativity which helps us understand partial charges

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