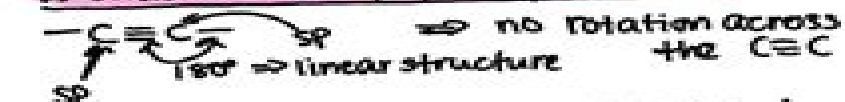


Alkynes

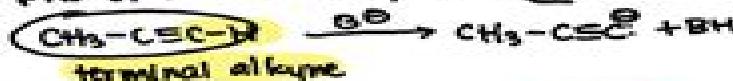
(A) Introduction:

1) Structure and physical properties:



- $\text{C}\equiv\text{C}$ formed of 1 σ bond and 2 π bond
- $\text{C}\equiv\text{C}$ triple bond adds 2 unsaturation
- formula: acyclic alkynes: C_{n+2}
- cyclic alkyne: C_{n+2m-4}

• pKa of terminal alkyne is 25.



• Internal alkyne: $\text{CH}_3-\text{C}\equiv\text{C}-\text{CH}_3$

• the melting point of alkynes increase with the # of carbons

* common names; terminal alkynes are named as substituted acetylenes

ex:



systematic: ethyne
 common: acetylene



1-butyne
 ethyl-
 acetylene



2-pentyne
 ethyl(methyl)-
 acetylene

$\text{HC}\equiv\text{C}-\text{CH}_2-\text{CH}_3$
 propargyl group

$\text{H}_2\text{C}\equiv\text{C}-\text{CH}_2-\text{CH}_3$
 allyl group

$\text{HC}\equiv\text{C}-\text{CH}_2-\text{Br}$
 propargyl bromide

(2) Nomenclature:

* same rules for alkenes but use suffix "yne" \Rightarrow alkyne

* multiple triple bonds:

2 = diynes; 3 = triynes; 4 = tetracynes



(3E,9S)-9-chlorodeca-3,9-diene
 stereo-
 chemical
 info

- lowest position for alkyne ($\# 1, 6$)
- alkene $\# 3$
- chiral carbonation L/E, R/S



⚠️ alcohols and amines have higher priority over alkenes and alkynes.

ex:



ex:



* above structures also show comparison between allyl and alkyne *

Ch 9 Alkynes Study Guide

Mary H. Bailey

Ch 9 Alkynes Study Guide:

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