

## Chapter 19

1. a) Contains as many hydrogen atoms as possible for the given number of carbons. Only carbon-carbon single bonds.

b) (i)  $C_{11}H_{24}$   
 (ii) liquid  
 (iii)  $C_{11}H_{24}(l) + 17O_2(g) \longrightarrow 11CO_2(g) + 12H_2O(l)$

2. a) a carbon-carbon double bond

b)

$$\begin{array}{c} \text{H} & \text{H} \\ | & | \\ \text{H}-\text{C} & -\text{C}=\text{C} \\ | & | \\ \text{H} & \text{H} \end{array}$$

c)

$$\begin{array}{c} \text{H} & \text{H} \\ | & | \\ \text{H}-\text{C} & -\text{C}=\text{C} \\ | & | \\ \text{H} & \text{H} \end{array} + Br_2 \longrightarrow \begin{array}{c} \text{H} & Br & \text{H} \\ | & | & | \\ \text{H}-\text{C} & -\text{C} & -\text{C}-Br \\ | & | & | \\ \text{H} & \text{H} & \text{H} \end{array}$$

(Home schooling parents: It doesn't matter what position you draw the bromine in on the right-hand carbon - IN, S or E (with hydrogens in the other two positions). All that matters is that one bromine becomes attached to each of the carbons either side of the original double bond.)

d) (i)  $CH_4(g) + Br_2(g) \longrightarrow CH_3Br(g) + HBr(g)$   
 (ii) In this (substitution) reaction, one of the hydrogens has been replaced by a bromine during the reaction. In the previous (addition) reaction, nothing was lost when the two molecules combined together.

3. a) Ethene is passed with steam at  $300^{\circ}\text{C}$  and 60 - 70 atmospheres over a phosphoric acid catalyst.  
 b) e.g. two off: continuous flow process, therefore faster; rapid reaction; purer product.  
 c) uses renewable resources; gentle conditions.

4. Dissolve the sugar in water in a conical flask (or similar) and add yeast. You now need some means of keeping the air out while allowing carbon dioxide to escape. You could use a commercial home-winemaking airlock of some sort, but a plug of cotton wool works just as well. Leave the mixture in the warm for a few days until bubbling stops.



Allow the yeast to settle, and pour off the solution so that it can be fractionally distilled, collecting the fraction which boils at  $78^{\circ}\text{C}$ . (See the diagram for the fractional distillation of an alcohol/water mixture in Chapter 11 of the book.)

continued . . .

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**Darrel Hess, Tom Lee  
McKnight, Stephen J. Stadler**

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