

**CHAPTER 6 REVIEW*****Chemical Bonding*****MIXED REVIEW****SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

**lattice energy** \_\_\_\_\_ a. ionic bond

**bond energy** \_\_\_\_\_ b. covalent bond

**enthalpy of vaporization** \_\_\_\_\_ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

**ionic** \_\_\_\_\_ a. H—F      **nonpolar covalent** \_\_\_\_\_ d. H—H

**ionic** \_\_\_\_\_ b. Na—Cl      **polar covalent** \_\_\_\_\_ e. H—C

**polar covalent** \_\_\_\_\_ c. H—O      **polar covalent** \_\_\_\_\_ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

**A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.**

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

**Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.**

**Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.**

# Covalent Bonding Guide Answers

**Shirley Soto**

## **Covalent Bonding Guide Answers:**

**A Teaching Guide to Revitalizing STEM Education** Daryao Khatri,2012-11-13 Standard math and science textbooks typically follow a deductive style of content presentation that involves too much lecturing too much of the teacher's back at the chalkboard too little interaction with students and too little time for all of the students to take adequate notes. By reading and using A Teaching Guide to Revitalizing STEM Education educators will rediscover how to streamline the subject matter math physics statistics and organic chemistry by eliminating unnecessary difficulties and distractions from course textbooks. A useful guide for both high school teachers and postsecondary faculty this book explains how to organize, arrange and streamline STEM content so that it is approachable, understandable and applicable for students. Likewise this guide discusses important classroom management skills and pedagogical techniques that will help students master these critical subjects. Providing and explaining over a dozen lesson plans A Teaching Guide to Revitalizing STEM Education will encourage educators to effectively optimize the recent emphases on science technology engineering and math education

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