

**Chemistry Unit 15: Chemical Equilibrium Worksheet****Practice 15.1 Writing Equilibrium Equations**Write the equilibrium expressions ( $K_{eq}$ ) for the following reactions:

1.  $2\text{CO(g)} \rightleftharpoons \text{C(s)} + \text{CO}_2\text{(g)}$
2.  $\text{N}_2\text{O}_4\text{(g)} \rightleftharpoons 2\text{NO}_2\text{(g)}$
3.  $2\text{HCl(g)} \rightleftharpoons \text{I}_2\text{(g)} + \text{Cl}_2\text{(g)}$
4.  $\text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)} \rightleftharpoons 2\text{Fe(s)} + 3\text{H}_2\text{O(g)}$
5.  $2\text{CO(g)} + 4\text{H}_2\text{O(g)} \rightleftharpoons 3\text{O}_2\text{(g)} + 2\text{CH}_4\text{(g)}$
6.  $\text{P}_4\text{(s)} + 6\text{NO(g)} \rightleftharpoons \text{P}_4\text{O}_6\text{(s)} + 3\text{N}_2\text{(g)}$

**15.2 Calculating Equilibrium Constants**Be sure to write out the equilibrium ( $K_{eq}$ ) expression for each problem before trying to solve it!

1. What is the equilibrium expression for the following reaction?  $\text{SeO}_3\text{(s)} + 2\text{CO(g)} \rightleftharpoons \text{Se(s)} + 2\text{CO}_2\text{(g)}$
2. At equilibrium the following concentrations are recorded: 20.0 M of  $\text{H}_2$ , 18.0 M of  $\text{CO}_2$ , 12.0 M  $\text{H}_2\text{O}$ , and 5.9 M of  $\text{CO}$  at 427°C. What is the value of  $K_{eq}$  at this temperature for the following reaction?  $\text{CO}_2\text{(g)} + \text{H}_2\text{(g)} \rightleftharpoons \text{CO(g)} + \text{H}_2\text{O(g)}$
3. At equilibrium a mixture of  $\text{N}_2$ ,  $\text{H}_2$ , and  $\text{NH}_3$  gas at 500°C is determined to consist of 0.602 M  $\text{N}_2$ , 0.420 M  $\text{H}_2$ , and 0.113 M of  $\text{NH}_3$ . What is the equilibrium constant at this temperature? Does this equilibrium favor the reactants or the products?  $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightleftharpoons 2\text{NH}_3\text{(g)}$
4. Given the equilibrium concentrations of 0.15 M  $\text{NO}$ , 0.12 M  $\text{H}_2$  and 1.6 M  $\text{N}_2$ . Calculate the equilibrium constant for the following reaction. Does this reaction favor the reactants or the products?  $2\text{NO(g)} + 2\text{H}_2\text{(g)} \rightleftharpoons \text{N}_2\text{(g)} + 2\text{H}_2\text{O(l)}$
5. What is the equilibrium constant for the following reaction if the equilibrium concentrations for the gases are 0.3 M for  $\text{C}_2\text{H}_6$ , 0.6 M for  $\text{O}_2$ , 1.3 M for  $\text{CO}_2$  and 1.8 M for  $\text{H}_2\text{O}$ ? Does the reaction favor the reactants or products?  $2\text{C}_2\text{H}_6\text{(g)} + 7\text{O}_2\text{(g)} \rightleftharpoons 4\text{CO}_2\text{(g)} + 6\text{H}_2\text{O(g)}$
6. Nitrogen dioxide reacts to form dinitrogen tetroxide. What is the equilibrium concentration of  $\text{N}_2\text{O}_4$ , if the  $K_{eq} = 4.70$  and the equilibrium concentration of  $\text{NO}_2$  is 0.0310 M?  $2\text{NO}_2\text{(g)} \rightleftharpoons \text{N}_2\text{O}_4\text{(g)}$

# Calculating Equilibrium Constant Chem Worksheet 18 3

**Sebastian Brünink**

### **Calculating Equilibrium Constant Chem Worksheet 18 3:**

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Sandler, 2017-04-24 In this newly revised 5th Edition of Chemical and Engineering Thermodynamics Sandler presents a modern applied approach to chemical thermodynamics and provides sufficient detail to develop a solid understanding of the key principles in the field The text confronts current information on environmental and safety issues and how chemical engineering principles apply in biochemical engineering bio technology polymers and solid state processing This book is appropriate for the undergraduate and graduate level courses

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Breneman, 1991 **Chemistry Homework** Frank Schaffer Publications, Joan DiStasio, 1996-03 Includes the periodic table writing formulas balancing equations stoichiometry problems and more

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of their main competitors Borland's Quattro Pro and Microsoft's Excel. Several compatible spreadsheets AsEasyAs ProQube might also be used as inexpensive alternatives. The author is a senior scientist at the Nuclear Research Centre in Karlsruhe Germany. He has been using spreadsheet software for nearly ten years successfully in a variety of chemical problems.

**Remediation Engineering** Suthan S. Suthersan, John Horst, Matthew Schnobrich, Nicklaus Welty, Jeff McDonough, 2016-11-25. This second edition of Remediation Engineering will continue to be the seminal handbook that regulators must have on hand to address any of the remediation issues they are grappling with daily. The book is wide ranging but specific enough to address any environmental remediation challenge. Patricia Reyes Interstate Technology Regulatory Council Washington DC USA. This book offers the researcher teacher practitioner student and regulator with state of the art advances in conducting site investigations and remediation for common and emerging contaminants. It is revolutionary in its approach to conducting subsurface investigation which greatly influences a successful and appropriate response in assessing and addressing environmental risk. This book is a giant leap forward in understanding how contaminants behave and how to reduce risk to acceptable levels in the natural world. Daniel T Rogers Amsted Industries Incorporated Chicago Illinois USA. This text is a superb reference and a good tool for learning about state of the art techniques in remediation of soil and groundwater. It will become a ready reference at many companies as the engineering community creates increased value from remediation efforts around the world. John Waites AVX Corporation Fountain Inn South Carolina USA. Remediation Engineering was first published in 1996 and quickly became the go to reference for a relatively young industry offering the first comprehensive look at the state of the science in treatment technologies of the time and the contaminants they applied to. This fully updated Second Edition will capture the fundamental advancements that have taken place during the last two decades within all the subdisciplines that form the foundation of the remediation engineering platform. It covers the entire spectrum of current technologies that are employed in the industry and also discusses future trends and how practitioners should anticipate and adapt to those needs. Features Shares the latest paradigms in remediation design approach and contaminant hydrogeology. Presents the landscape of new and emerging contaminants. Details the current state of the practice for both conventional technologies such as sparging and venting. Examines newer technologies such as dynamic groundwater recirculation and injection based remedies to address both organic and inorganic contaminants. Describes the advances in site characterization concepts such as smart investigations and digital conceptual site models. Includes all new color photographs and figures.

**INIS Atomindex**, 1984

**Geochemistry International**, 1998. Vols for 1964 v 2 no 1 1965 include selected articles translated from geochemical papers from other languages but primarily from Russian German French and Japanese.

*Tutorial, Human Factors in Software Development* Bill Curtis, 1981. Presents a broad overview of the experimental research on human factors in software development.

[Data Sources](#), 2000

**Study Guide for Chemical Principles** Thomas Elliott Taylor, 1979

Equilibrium Calculations John T. Donoghue,1971      *Chemical Equilibrium* William Guenther,1975-12 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

**Chemical Equilibrium In a Nutshell** Brian Coleman,2018-06-25 Understanding the math and minutiae of chemical equilibrium can be a tall task for anyone so why not enlist the help of a scientific squirrel to guide you on your journey Join Dr Wash as we dabble in equilibrium constants and other tools needed to predict chemical processes This book focuses on introductory concepts at the high school and early university level focusing on identifying equilibrium calculating K and Q discussing Le Chatelier's principle and tying equilibrium with the field of Thermodynamics Full of step by step instructions and practice questions this book aims to simplify one of the more complex topics found within the field of chemistry

Study Guide [to Accompany] General Chemistry James E. Brady,1982      **Equilibrium** Thomas R. Blackburn,1969      **The Computation of Chemical**

**Equilibria** F. van Zeggeren,S. H. Storey,2011-02-17 This 1970 book the authors derive the equations describing equilibria in different types of system and outline the effect of variation of the parameters of the system on the equilibrium composition by using equilibrium calculations in high temperature high pressure processes in rocketry and in explosives technology

Chemical Calculations Paul Yates,1997-06-05 Chemical Calculations provides an introduction to the mathematics required for physical chemistry courses This book is unique in that it provides a gentle introduction with a chemistry centered rather than math centered approach Written by a chemist for undergraduate students it imparts an understanding of the subject from a chemist's viewpoint using examples from real chemistry It includes illustrations that show exactly how to use calculators to work problems and examples of important chemical problems with fully worked solutions This book is an ideal companion throughout a chemistry course that can be consulted when required and used to keep one step ahead of the lecture

Chemical Equilibria Harry L. Pardue,2018-11-15 Concepts procedures and programs described in this book make

it possible for readers to solve both simple and complex equilibria problems quickly and easily and to visualize results in both numerical and graphical forms They allow the user to calculate concentrations of reactants and products for both simple and complicated situations The user can spend less time doing calculations and more time thinking about what the results mean in terms of a larger problem in which she or he may be interested

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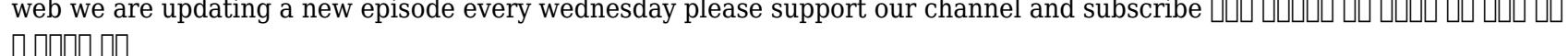
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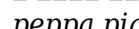
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