

Solution Stoichiometry

1. What volume of a 0.35M AgNO_3 is required to completely react with 25 mL of a 0.24M NaCl solution?



Required: V of AgNO_3

Start with the other reactant

$$\frac{0.24 \text{ mol NaCl}}{1 \text{ L}} \cdot \frac{0.055 \text{ L NaCl}}{1 \text{ mol NaCl}} \cdot \frac{1 \text{ mol AgNO}_3}{1 \text{ mol NaCl}}$$

These cancel

$$\frac{0.35 \text{ mol AgNO}_3}{0.35 \text{ mol AgNO}_3} = 0.037 \text{ L} = 37.7 \text{ mL}$$

2. 37 mL of a 0.24M KI solution is required to completely react with 42 mL of a $\text{Pb}(\text{NO}_3)_2$ solution. What was the original concentration of the lead (II) Nitrate solution?



Given: $\text{V}_1 = 37 \text{ mL KI}$ cm^3
 $\text{V}_2 = 42 \text{ mL Pb}(\text{NO}_3)_2$
 $\text{C}_1 = 0.24 \text{ M}$ KNO_3

Required: original concentration of the lead

$$\frac{0.24 \text{ mol KI}}{1 \text{ L}} \cdot \frac{0.037 \text{ L}}{1} \cdot \frac{1 \text{ mol Pb}(\text{NO}_3)_2}{2 \text{ mol KI}} \cdot \frac{1}{0.037 \text{ L}} = 0.1277$$

$$2(42) \text{ M}_2 = 1(0.24)(37)$$

$$\text{M}_2 = 10.72 / 84$$

$$\text{M}_2 = 0.1277 \text{ M}$$

3. What mass of solid Magnesium Hydroxide can be produced if 15 mL of a 0.63M $\text{Mg}(\text{NO}_3)_2$ solution reacts completely with excess NaOH ?



Required: mass of solid Magnesium Hydroxide if 15 mL of a 0.63M $\text{Mg}(\text{NO}_3)_2$

Given: 15 mL $\text{Mg}(\text{NO}_3)_2$, 0.63M $\text{Mg}(\text{NO}_3)_2$

$$\frac{1 \text{ mol AgNO}_3}{1 \text{ mol NaCl}} \cdot \frac{0.63 \text{ mol Mg}(\text{NO}_3)_2}{1 \text{ mol NaCl}} \cdot \frac{0.015 \text{ L}}{1 \text{ L}} \cdot \frac{1 \text{ mol Mg(OH)}_2}{1 \text{ mol Mg}(\text{NO}_3)_2}$$

$$\cdot \frac{58.32 \text{ g}}{1 \text{ mol Mg(OH)}_2} = 1.65 \text{ g Mg(OH)}_2$$

4. What mass of solid Calcium Phosphate can be produced if 78mL of a 0.17M Sodium Phosphate solution is mixed with 58mL of a 0.21M Calcium Chloride solution?



Given:

$$\frac{78 \text{ mL Na}_3\text{PO}_4}{0.17 \text{ M Na}_3\text{PO}_4} \cdot \frac{58 \text{ mL CaCl}_2}{0.21 \text{ M CaCl}_2}$$

$$\text{molar mass:} \\ \text{Ca} = 2 \cdot 40.08 \\ \text{P} = 2 \cdot 30.97 \\ \text{O} = 2 \cdot 16 \text{ mol}$$

Required: mass(g) $\text{Ca}_3(\text{PO}_4)_2$

$$\frac{0.17 \text{ mol Na}_3\text{PO}_4}{1 \text{ L}} \cdot \frac{0.078 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol Na}_3\text{PO}_4}$$

$$\cdot \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 1.925 \text{ g Ca}_3(\text{PO}_4)_2$$

From CaCl_2

$$\frac{0.21 \text{ mol}}{1 \text{ L}} \cdot \frac{0.058 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol CaCl}_2} \cdot \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2}$$

$$= 1.4371 \text{ g Ca}_3(\text{PO}_4)_2$$

Calculating With Stoichiometry Student Guide Answers

Pamela Mills,Amina El-Ashmawy

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