

Depo-Provera Perpetual Calendar

4-TIMES-A-YEAR DOSING FLEXIBILITY

[based on 3-month (13-week) dosing intervals, with the flexibility of dosing between weeks 11 and 13]

GIVEN	DUE
Jan 1	Mar 19-Apr 2
Jan 2	Mar 20-Apr 3
Jan 3	Mar 21-Apr 4
Jan 4	Mar 22-Apr 5
Jan 5	Mar 23-Apr 6
Jan 6	Mar 24-Apr 7
Jan 7	Mar 25-Apr 8
Jan 8	Mar 26-Apr 9
Jan 9	Mar 27-Apr 10
Jan 10	Mar 28-Apr 11
Jan 11	Mar 29-Apr 12
Jan 12	Mar 30-Apr 13
Jan 13	Mar 31-Apr 14
Jan 14	Apr 1-Apr 15
Jan 15	Apr 2-Apr 16
Jan 16	Apr 3-Apr 17
Jan 17	Apr 4-Apr 18
Jan 18	Apr 5-Apr 19
Jan 19	Apr 6-Apr 20
Jan 20	Apr 7-Apr 21
Jan 21	Apr 8-Apr 22
Jan 22	Apr 9-Apr 23
Jan 23	Apr 10-Apr 24
Jan 24	Apr 11-Apr 25
Jan 25	Apr 12-Apr 26
Jan 26	Apr 13-Apr 27
Jan 27	Apr 14-Apr 28
Jan 28	Apr 15-Apr 29
Jan 29	Apr 16-Apr 30
Jan 30	Apr 17-May 1
Jan 31	Apr 18-May 2
Feb 1	Apr 19-May 3
Feb 2	Apr 20-May 4
Feb 3	Apr 21-May 5
Feb 4	Apr 22-May 6
Feb 5	Apr 23-May 7
Feb 6	Apr 24-May 8
Feb 7	Apr 25-May 9
Feb 8	Apr 26-May 10
Feb 9	Apr 27-May 11
Feb 10	Apr 28-May 12
Feb 11	Apr 29-May 13
Feb 12	Apr 30-May 14
Feb 13	May 1-May 15
Feb 14	May 2-May 16
Feb 15	May 3-May 17

GIVEN	DUE
Feb 16	May 4-May 18
Feb 17	May 5-May 19
Feb 18	May 6-May 20
Feb 19	May 7-May 21
Feb 20	May 8-May 22
Feb 21	May 9-May 23
Feb 22	May 10-May 24
Feb 23	May 11-May 25
Feb 24	May 12-May 26
Feb 25	May 13-May 27
Feb 26	May 14-May 28
Feb 27	May 15-May 29
Feb 28	May 16-May 30
Mar 1	May 17-May 31
Mar 2	May 18-Jun 1
Mar 3	May 19-Jun 2
Mar 4	May 20-Jun 3
Mar 5	May 21-Jun 4
Mar 6	May 22-Jun 5
Mar 7	May 23-Jun 6
Mar 8	May 24-Jun 7
Mar 9	May 25-Jun 8
Mar 10	May 26-Jun 9
Mar 11	May 27-Jun 10
Mar 12	May 28-Jun 11
Mar 13	May 29-Jun 12
Mar 14	May 30-Jun 13
Mar 15	May 31-Jun 14
Mar 16	Jun 1-Jun 15
Mar 17	Jun 2-Jun 16
Mar 18	Jun 3-Jun 17
Mar 19	Jun 4-Jun 18
Mar 20	Jun 5-Jun 19
Mar 21	Jun 6-Jun 20
Mar 22	Jun 7-Jun 21
Mar 23	Jun 8-Jun 22
Mar 24	Jun 9-Jun 23
Mar 25	Jun 10-Jun 24
Mar 26	Jun 11-Jun 25
Mar 27	Jun 12-Jun 26
Mar 28	Jun 13-Jun 27
Mar 29	Jun 14-Jun 28
Mar 30	Jun 15-Jun 29
Mar 31	Jun 16-Jun 30
Apr 1	Jun 17-Jul 1
Apr 2	Jun 18-Jul 2

GIVEN	DUE
Apr 3	Jun 19-Jul 3
Apr 4	Jun 20-Jul 4
Apr 5	Jun 21-Jul 5
Apr 6	Jun 22-Jul 6
Apr 7	Jun 23-Jul 7
Apr 8	Jun 24-Jul 8
Apr 9	Jun 25-Jul 9
Apr 10	Jun 26-Jul 10
Apr 11	Jun 27-Jul 11
Apr 12	Jun 28-Jul 12
Apr 13	Jun 29-Jul 13
Apr 14	Jun 30-Jul 14
Apr 15	Jul 1-Jul 15
Apr 16	Jul 2-Jul 16
Apr 17	Jul 3-Jul 17
Apr 18	Jul 4-Jul 18
Apr 19	Jul 5-Jul 19
Apr 20	Jul 6-Jul 20
Apr 21	Jul 7-Jul 21
Apr 22	Jul 8-Jul 22
Apr 23	Jul 9-Jul 23
Apr 24	Jul 10-Jul 24
Apr 25	Jul 11-Jul 25
Apr 26	Jul 12-Jul 26
Apr 27	Jul 13-Jul 27
Apr 28	Jul 14-Jul 28
Apr 29	Jul 15-Jul 29
Apr 30	Jul 16-Jul 30
May 1	Jul 17-Jul 31
May 2	Jul 18-Aug 1
May 3	Jul 19-Aug 2
May 4	Jul 20-Aug 3
May 5	Jul 21-Aug 4
May 6	Jul 22-Aug 5
May 7	Jul 23-Aug 6
May 8	Jul 24-Aug 7
May 9	Jul 25-Aug 8
May 10	Jul 26-Aug 9
May 11	Jul 27-Aug 10
May 12	Jul 28-Aug 11
May 13	Jul 29-Aug 12
May 14	Jul 30-Aug 13
May 15	Jul 31-Aug 14
May 16	Aug 1-Aug 15
May 17	Aug 2-Aug 16
May 18	Aug 3-Aug 17

GIVEN	DUE
May 19	Aug 4-Aug 18
May 20	Aug 5-Aug 19
May 21	Aug 6-Aug 20
May 22	Aug 7-Aug 21
May 23	Aug 8-Aug 22
May 24	Aug 9-Aug 23
May 25	Aug 10-Aug 24
May 26	Aug 11-Aug 25
May 27	Aug 12-Aug 26
May 28	Aug 13-Aug 27
May 29	Aug 14-Aug 28
May 30	Aug 15-Aug 29
May 31	Aug 16-Aug 30
Jun 1	Aug 17-Aug 31
Jun 2	Aug 18-Sept 1
Jun 3	Aug 19-Sept 2
Jun 4	Aug 20-Sept 3
Jun 5	Aug 21-Sept 4
Jun 6	Aug 22-Sept 5
Jun 7	Aug 23-Sept 6
Jun 8	Aug 24-Sept 7
Jun 9	Aug 25-Sept 8
Jun 10	Aug 26-Sept 9
Jun 11	Aug 27-Sept 10
Jun 12	Aug 28-Sept 11
Jun 13	Aug 29-Sept 12
Jun 14	Aug 30-Sept 13
Jun 15	Aug 31-Sept 14
Jun 16	Sept 1-Sept 15
Jun 17	Sept 2-Sept 16
Jun 18	Sept 3-Sept 17
Jun 19	Sept 4-Sept 18
Jun 20	Sept 5-Sept 19
Jun 21	Sept 6-Sept 20
Jun 22	Sept 7-Sept 21
Jun 23	Sept 8-Sept 22
Jun 24	Sept 9-Sept 23
Jun 25	Sept 10-Sept 24
Jun 26	Sept 11-Sept 25
Jun 27	Sept 12-Sept 26
Jun 28	Sept 13-Sept 27
Jun 29	Sept 14-Sept 28
Jun 30	Sept 15-Sept 29
Jul 1	Sept 16-Sept 30
Jul 2	Sept 17-Oct 1
Jul 3	Sept 18-Oct 2

Long-acting, Reversible

Depo-Provera
Contraceptive Injection

medroxyprogesterone acetate injectable suspension

4 Times a Year

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A Woman's Guide to Sexual Health Mary Jane Minkin, Carol V. Wright, 2005-01-01 A prominent gynecologist explains how to make the best choices for female reproductive health in this authoritative easy to read guide This book is for every woman who has wished for an unhurried personal conversation with a sympathetic doctor who will answer her questions about reproductive health Dr Mary Jane Minkin a gynecologist practicing for more than 25 years presents a complete and up to date guide to a healthy reproductive system for women in their teens through middle age With warmth and understanding Dr Minkin and coauthor Carol V Wright respond to questions about the gynecological issues that concern women today including sexual activity contraception and family planning Readers of The Yale Guide to Women s Reproductive Health will learn how the female body works what problems may arise and what solutions are available in short they will become better prepared to participate in their own health care and to make healthy decisions

Pharmacology Joyce LeFever Kee, Evelyn R. Hayes, Linda E. McCuistion, 2014-01-30 Previous edition has subtitle a nursing process approach

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2 \text{Na}^+ \text{aq}$ $2 \text{Cl}^- \text{aq}$ $\text{Cu}^{2+} \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $2 \text{Na}^+ \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $\text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+ \text{aq}$ $\text{Cl}^- \text{aq}$ $\text{Na}^+ \text{aq}$ $\text{HCO}_3^- \text{aq}$ $\text{Na}^+ \text{aq}$ $\text{Cl}^- \text{aq}$ H_2O $\text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+ net ionic equation $\text{H}^+ \text{aq}$ $\text{Cl}^- \text{aq}$ $\text{Na}^+ \text{aq}$ $\text{HCO}_3^- \text{aq}$ $\text{Na}^+ \text{aq}$ $\text{Cl}^- \text{aq}$ H_2O $\text{CO}_2 \text{g}$ $\text{H}^+ \text{aq}$ $\text{HCO}_3^- \text{aq}$ H_2O $\text{CO}_2 \text{g}$ here is a video on spectator ions *chapter 8 04 ionic equations chemistry libretexts* - Mar 06 2023

web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3 \text{Ba}^{2+} \text{aq}$ $2 \text{PO}_4^{3-} \text{aq}$ $\text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3 \text{aq} + \text{KCl aq} \rightarrow \text{AgCl s} + \text{KNO}_3 \text{aq}$ total ionic $\text{Ag}^+ \text{aq} + \text{NO}_3^- \text{aq} + \text{K}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl s} + \text{K}^+ \text{aq} + \text{NO}_3^- \text{aq}$ net ionic $\text{Ag}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl s}$

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O l}$ the spectator ions in KOH aq and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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web ionic equations and spectator ions in chemistry an ion is an atom or molecule that has a net electric charge due to the loss or gain of one or more electrons ions containing positive charge are cations and those containing negative charge are anions an ion with a net charge of zero is electrically neutral

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web apr 6 2022 spectator ion a spectator ion is an ion that exists as a reactant and a product in a chemical equation a spectator ion can therefore be observed in the reaction of aqueous solutions of sodium carbonate and copper ii sulfate but does not affect the equilibrium what are spectator ions

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web 18 terms lourdes l22 preview terms in this set 4 molecular equation showing the complete chemical formulas of reactions and products $\text{Pb(NO}_3)_2 \text{aq} + 2\text{Kl aq} \rightarrow \text{PbI}_2 \text{s} + 2\text{KNO}_3 \text{aq}$ complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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web the Ag^+ and Cl^- ions become AgCl s but the Na^+ ions and the NO_3^- ions stay as Na^+ ions and NO_3^- ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jul 26 2017 answer the spectator ions are explanation in the net ionic equations we are not include the spectator ions in the equations spectator ions the ions present on reactant and product side 16 used to grind chemicals to powder tmraor nda stlepe star 4 4 5 heart 181 verified verified answer food

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web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction the questions are intended to test the students understanding of the meaning of the reactions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{Cl}^- + \text{Cu}^{2+} \rightarrow \text{CuCl}_2 \text{ s}$ $2\text{Cl}^- + \text{Cu}^{2+} \rightarrow \text{CuCl}_2 \text{ s}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length