

# SCIENCE 8 – DENSITY CALCULATIONS WORKSHEET

NAME: \_\_\_\_\_

- 1) A student measures the mass of an 8 cm<sup>3</sup> block of brown sugar to be 12.9 g. What is the density of the brown sugar?

$$D = \frac{M}{V} = \frac{12.9g}{8cm^3} = 1.6125 = \boxed{1.61g/cm^3}$$

- 2) A chef fills a 50 mL container with 43.5 g of cooking oil. What is the density of the oil?

$$D = \frac{M}{V} = \frac{43.5g}{50mL} = \boxed{0.87g/mL}$$

- 3) Calculate the mass of a liquid with a density of 2.5 g/mL and a volume of 15 mL.

$$\frac{15mL}{1} \times \frac{2.5g}{1mL} = \boxed{37.5g}$$

- 4) Calculate the volume of a liquid with a density of 5.45 g/mL and a mass of 65 g.

$$\frac{65g}{1} \times \frac{1mL}{5.45g} = \boxed{11.9mL}$$

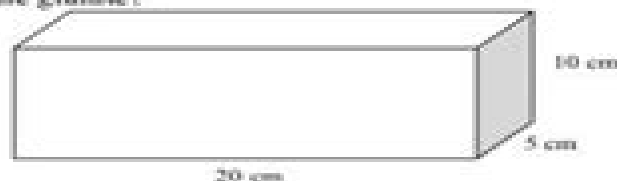
- 5) A machine shop worker records the mass of an aluminum cube as 176 g. If one side of the cube measures 4 cm, what is the density of the aluminum?

$$V = 4cm \times 4cm \times 4cm = 64cm^3 \quad D = \frac{176g}{64cm^3} = \boxed{2.75g/cm^3}$$

- 6) A teacher performing a demonstration finds that a piece of cork displaces 23.5 mL of water. The piece of cork has a mass of 5.7 g. What is the density of the cork?

$$D = \frac{M}{V} = \frac{5.7g}{23.5mL} = 0.2425531915 = \boxed{0.243g/mL}$$

- 7) A carver begins work on the following block of granite that weighs 2700 g. What is the density of the granite?



$$V = 20cm \times 5cm \times 10cm$$

$$V = 1,000 cm^3$$

$$D = \frac{2700g}{1000cm^3} = \boxed{2.7g/cm^3}$$

- 8) A piece of PVC plumbing pipe displaces 60 mL when placed into a container of water. If the pipe has a mass of 78 g, what is the density of PVC?

$$D = \frac{M}{V} = \frac{78g}{60mL} = \boxed{1.3g/mL}$$

- 9) A solid magnesium flare has a mass of 1300 g and a volume of 743 cm<sup>3</sup>. What is the density of the magnesium?

$$D = \frac{M}{V} = \frac{1300g}{743cm^3} = 1.749663526 = \boxed{1.75g/cm^3}$$

# Density Practice Problems With Answers Fordhamprep

**Hussin A.Rothana**



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web  $m \times c \times \Delta t$  equation 1 where  $m$  is the total mass of the solution solute plus solvent  $c$  is the specific heat of the solution and  $\Delta t$  is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of  
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web the heat of reaction  $\Delta H_{\text{soln}}$  is written after the products in units of  $\text{kJ mol rxn}$  for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as  $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$   $\Delta H_{\text{soln}} = -44.2 \text{ kJ mol rxn}$  write equations to similarly represent the dissolving process for

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