



# COVALENT BONDING

## Practice Problems

Do your best to do better than your previous problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

1. Draw a molecular formula for each of the following compounds.

- a. Silicon tetrachloride      b. Nitrogen dioxide      c. Hydrogen peroxide  
d. Sulfur hexafluoride      e. Dinitrogen pentoxide

2. Write the formula for the following molecular compounds.

- a. Carbon monoxide      b. Diphosphorus pentoxide      c. Nitrogen  
d. Sulfur      e. Hydrogen      f. Oxygen

3. Draw a Lewis structure for each of the following molecular compounds. Indicate the type of hybridization for each atom.

4. Draw a Lewis structure for each of the following molecular compounds. Indicate the type of hybridization for each atom. Indicate the type of hybridization for each atom.

5. Draw a Lewis structure for each of the following molecular compounds. Indicate the type of hybridization for each atom. Indicate the type of hybridization for each atom.

### SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw a Lewis structure for each of the following compounds.



2. Draw a Lewis structure for each of the following compounds.



3. Draw a Lewis structure for each of the following compounds.



4. Draw a Lewis structure for each of the following compounds.

5. Draw a Lewis structure for each of the following compounds. Indicate the type of hybridization for each atom.



Indicate the type of hybridization for each atom.

### SECTION 8.3 BONDING IN THE LIGANDS

1. Draw a Lewis structure for each of the following ligands. Indicate the type of hybridization for each atom.

2. Draw a Lewis structure for each of the following ligands. Indicate the type of hybridization for each atom.

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**E Staff**



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