

# COVALENT BONDING

## Practice Problems

Use your knowledge of covalent bonding to answer the following questions.

### SOLVING PROBLEMS INVOLVING COVALENT COMPOUNDS

1. Identify the following covalent compounds:

a.  $\text{H}_2\text{O}$       b.  $\text{NH}_3$       c.  $\text{CH}_4$

d.  $\text{H}_2\text{S}$       e.  $\text{H}_2\text{O}_2$

f.  $\text{H}_2\text{O}_2$       g.  $\text{H}_2\text{O}_2$

h.  $\text{H}_2\text{O}_2$       i.  $\text{H}_2\text{O}_2$

2. Identify the following covalent compounds as polar or nonpolar molecules:

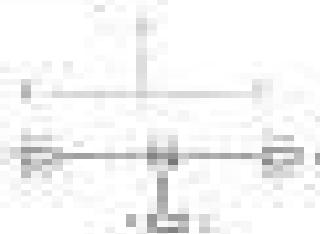
a.  $\text{H}_2\text{O}$       b.  $\text{NH}_3$       c.  $\text{CH}_4$       d.  $\text{H}_2\text{S}$       e.  $\text{H}_2\text{O}_2$

f.  $\text{H}_2\text{O}_2$       g.  $\text{H}_2\text{O}_2$       h.  $\text{H}_2\text{O}_2$       i.  $\text{H}_2\text{O}_2$

### SOLVING PROBLEMS: THE LANGUAGE OF COVALENT BONDING



a. Draw the Lewis structure for the molecule.



b. Draw the Lewis structure for the molecule.



c. Draw the Lewis structure for the molecule.

d. Draw the Lewis structure for the molecule.

### SOLVING PROBLEMS: IDENTIFYING THE COMPOUNDS

a. Which element has the greater electron affinity?  $\text{O}$  or  $\text{N}$

b. Which element has the greater electron affinity?  $\text{O}$  or  $\text{N}$

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