

the Atom

BASIC BUILDING BLOCK OF MATTER

CANNOT BE BROKEN DOWN

ORBITAL

- circular paths of energy
- 300 known as energy levels

NEUTRON

- 1 amu
- NO charge (neutral)
- located in nucleus

ATOMIC MASS = PROTONS + NEUTRONS

ISOTOPE

atoms of the same element with different numbers of neutrons
e.g. carbon - 12

ATOMIC MASS

- mass, in amu, of 1 mole of an element

ATOMIC NUMBER

- identifies an element
- number of protons
- number of electrons

ELECTRON CONFIGURATION

- amount of electrons in each energy level

All atoms
are neutral

ELECTRON

- 0 amu
- -1 charge
- located outside nucleus, in orbitals

PROTON

- 1 amu
- +1 charge
- located in nucleus

SUM MOLES

• mole of substance

AVOGADRO'S NUMBER

* 1 mole = 6.02×10^{23} atoms *

Molar mass
(mass atomic mass)

- mass of one mole of a substance

OXIDATION STATES

- charge an element can take on while in a compound

SUBSCRIPT

12	→ mass number
C	→ element symbol
6	→ atomic number

carbon - 12

ELEMENT SYMBOL
• condensed form of an element's name

$$\text{AVERAGE ATOMIC MASS} = (\text{mass of isotope 1} \times \frac{\text{no. of isotopes}}{100}) + (\text{mass of isotope 2} \times \frac{\text{no. of isotopes}}{100}) + \dots$$

Defining The Atom Study Guide

Arshad Iqbal

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