

# the Atom

**BASIC BUILDING BLOCK OF MATTER**

**CANNOT BE BROKEN DOWN**

## ORBITAL

- Circular paths of energy
- also known as energy levels

## NEUTRON

- 1 amu
- No charge (neutral)
- located in nucleus

**MASS NUMBER = Protons + Neutrons**

## Isotope

atoms of the same element  
with different numbers of protons  
e.g. carbon-12

## ATOMIC MASS

- mass, in g/mol, of 1 mole of an element

## ATOMIC NUMBER

- identifies an element
- Number of protons
- Number of electrons

## ELECTRON CONFIGURATION

- amount of electrons in each energy level

## ELEMENT SYMBOL

- condensed form of an element's name

## ELECTRON

- 0 amu
- -1 charge
- located outside nucleus, in orbitals

## PROTON

- 1 amu
- +1 charge
- located in nucleus

## MOLE

- amount of substance

## AVOGADRO'S NUMBER

\* 1 mole =  $6.02 \times 10^{23}$  atoms \*

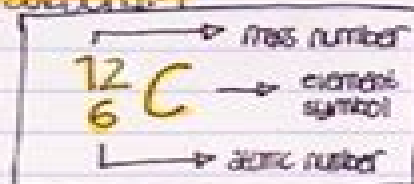
**molar mass**  
(atomic mass units)

- mass of one mole of a substance

## OXIDATION STATES

- charge an element can take on while in a compound

## SUBSCRIPT



carbon - 12

All atoms are neutral

$$\text{AVERAGE ATOMIC MASS} = (\text{mass of isotope 1} \times \frac{\%}{100}) + (\text{mass of isotope 2} \times \frac{\%}{100}) + \dots$$

# Defining The Atom Study Guide

**Arshad Iqbal**



## Defining The Atom Study Guide:

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