

Gas Laws Study Guide

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Name: _____ Period: _____ Date: _____

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Initial Temperature: T_1 in $^{\circ}\text{C}$ or K
Initial Pressure: P_1 in atm , mm Hg , in Hg , Pa , psi , torr
Initial Volume: V_1 in L or m^3

Match the following to the equations:

$\frac{P_1}{T_1} = \frac{P_2}{T_2}$

Standardized Gas Law

$P_1V_1 = P_2V_2$

Boyle's Law

$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

Ideal Gas Law

$PV = nRT$

Charles's Law

$\frac{P_1}{P_2} = \frac{V_1}{V_2}$

Gay-Lussac's Law

Compare the following statements to the equations above to determine the correct answer.
a) Pressure increases directly proportional to temperature given that _____ increases resulting from Boyle's Law or Charles's Law.

b) Pressure increases directly proportional to temperature given that _____ increases resulting from Boyle's Law or Charles's Law.

c) The number of moles of gas in a container increases the greater increase in the pressure given that _____ increases, and the temperature increases _____.

Chemistry Study Guide Gas Laws

Paul B. Kelter

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Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text. The new edition of the Study Guide includes many new worked out examples. *Essential Concepts of Chemistry Study Guide* James R. Braun, Sherman, 1999

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change *Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application. **Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter, 1993 *Chemistry Super Review - 2nd Ed.* Editors of

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