

## 14: Chemical Bonding Theories

## Key Chemistry Terms

- Ionic Bond:** Bond formed from electrostatic attraction between ions (charged atoms)—formed from metals with nonmetals.
- Covalent Bond:** A bond formed between two nonmetals that involves shared electrons.
- Polar covalent bond:** Bond formed between two nonmetals—uneven sharing of electrons.
- Metallic bond:** Formed between metal atoms—electrons are pooled between the network of atoms.
- Polar bond:** Bond between two atoms with a great difference in electronegativities—uneven sharing of electrons resulting in a slightly positive region and a slightly negative region of the bond.
- Electronegativity:** An atom's "pull" on electrons shared with another atom.
- Polar molecule:** Molecule where the polar bonds do not cancel each other out in 3D orientation.
- Isomers:** Molecules with the same chemical formula, but different bonding structures.
- Resonance:** Two compounds with multiple bonds in different locations.
- Valence Bond Theory:** Overlap of atomic orbitals form bond.
- Sigma ( $\sigma$ ) bonds:** First bond between two atoms. Formed from head-on overlap of orbitals.
- Pi ( $\pi$ ) bonds:**  $2^{\text{nd}}$  or  $3^{\text{rd}}$  bond between two atoms. Formed from overlap of parallel p orbitals.
- Hybridization:** Atomic orbitals hybridize into orbitals of the same energy for all sigma bonds of the atom.

## Characteristics of bond types

Bond type	Partners	Electrons are
Ionic	Metal & non-metal	Transferred
Covalent	Non-metals	Shared
Polar Covalent	Non-metals	Unevenly shared
Metallic	Metals	Pooled

## Common characteristics:

- Ionic:** High melting points, most dissolve in water, conduct electricity when dissolved in water, brittle.
- Covalent:** Low melting points, most do not dissolve in water, do not conduct electricity when dissolved in water.
- Polar covalent:** Medium melting points, some dissolve in water, do not conduct electricity when dissolved in water.
- Metallic:** Soft, conduct heat and electricity, do not dissolve in water.

## Bond Polarity

When nonmetals bond covalently with a large difference in electronegativity.

- Attribute value of difference:
  - $0 - 0.4 = \text{covalent}$
  - $0.5 - 1.4 = \text{polar covalent}$
  - $1.5 - 4 = \text{ionic}$

Symbolized with an arrow pointing towards the more electronegative element and a crossed tail by the less electronegative element.

Example:

$\text{C} - \text{H}$  Electronegativity difference of 0.5 = polar bond



## Isomers and Resonance

## Isomer examples:



## Resonance examples:



## Sigma and Pi bonds

How many sigma and pi bonds are in the following?



6 sigma bonds & 3 pi bonds

## Hybridization

# of sigma bonds	Hybridization	Formed from
2	$\text{sp}^3$	1 s and 3 p
3	$\text{sp}^2$	1 s and 2 p
4	$\text{sp}^2$	1 s and 3 p

# Chemistry The Physical Setting 2014 Answer Key

**Barry A. Wills,James Finch**

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of the issues faced in all aspects of drilling operations with a particular focus on the unique aspects of offshore operations

*Ecological Solutions to Agricultural Land Degradation* Subhash Babu, Anup Das, Sanjay Singh Rathore, Raghavendra Singh, Rajbir Singh, 2025-05-20 This book provides an overview of the current status of land degradation globally and the different approaches used for its restoration. In addition to the technical aspects it also delves into the social and economic implications of land restoration such as in terms of policy planning. Land degradation is a serious threat to agriculture affecting soil function and productivity. There are up to 6 billion hectares of degraded soil worldwide. It is estimated that 30% of forests, 20% of agricultural land and 10% of rangelands are severely affected by land degradation. The main causes of land degradation are soil erosion, acidification, compaction and salinization, intensive agriculture, overgrazing, poor management of arable land and deforestation. The consequences of such degradation include sand dune expansion and displacement, reduced rainfall, rangeland depletion, wind and rain erosion, and loss of biomass and biological integrity. In 2015 the global goal of achieving land degradation neutrality (LDN) was included in the 2030 Agenda for Sustainable Development as sustainable development goal 15.3. Developing productive soil ecosystems by restoring degraded soils is therefore a future perspective for food security and environmental sustainability. This can be achieved through crop rotation, agroforestry, conservation agriculture, cover crops, terrace farming, ditching, geo jute and contour farming, and zero tillage. Ecological restoration of farmlands offers an opportunity to reconcile agricultural production with improving ecosystem services. Restoring degraded land through environmentally friendly agroecological approaches has the potential to improve agricultural production and provide healthy ecosystems for organisms. This book is an important resource for students, academics, researchers, national and international agricultural scientists and soil experts.

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