

## Unit 2 Worksheet 3 – PVTn Problems

On each of the problems below, start with the given P, V, T, or n; then make a decision as to how a change in P, V, T, or n will affect the starting quantity, and then multiply by the appropriate factor. Draw particle diagrams of the initial and final conditions.

1. A sample of gas occupies 150 mL at 25 °C. What is its volume when the temperature is increased to 50 °C? (P and n = constant)

	P	T	V	n
Initial	—	298K	150mL	—
Final	—	323K	?	—
Effect	—	↑	↑	—

$$V_2 = 150\text{mL} \times \frac{323\text{K}}{298\text{K}} = 163\text{mL} \rightarrow 160\text{mL}$$

2. The pressure in a bicycle tire is 105 psi at 25 °C here in Fresno. You take the bicycle up to Huntington, where the temperature is – 5 °C. What is the pressure in the tire? (V and n = constant)

	P	T	V	n
Initial	105psi	298K	—	—
Final	?	268K	—	—
Effect	↓	↓	—	—

$$P_2 = 105\text{psi} \times \frac{268\text{K}}{298\text{K}} = 94.4\text{psi}$$

3. What would be the new pressure if 250 cm<sup>3</sup> of gas at standard pressure is compressed to a volume of 150 cm<sup>3</sup>? (T, n = constant)

	P	T	V	n
Initial	1atm	—	250cm <sup>3</sup>	—
Final	?	—	150cm <sup>3</sup>	—
Effect	↑	—	↓	—

$$P_2 = 1\text{atm} \times \frac{250\text{cm}^3}{150\text{cm}^3} = 1.7\text{atm}$$

or 1300mmHg  
or 170kPa  
or 25psi

# Chemistry Unit 2 Pvtn Problems Answers

**Ian Pickup**



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