

## Unit 2 Worksheet 3 – PVTn Problems

On each of the problems below, start with the given P, V, T, or n; then make a decision as to how a change in P, V, T, or n will affect the starting quantity, and then multiply by the appropriate factor. Draw particle diagrams of the initial and final conditions.

1. A sample of gas occupies 150 mL at 25 °C. What is its volume when the temperature is increased to 50°C? (P and n = constant)

	P	T	V	n
Initial	--	298K	150mL	--
Final	--	323K	?	--
Effect	--	↑	↑	--

$$P_2 = 150 \text{ mL} \times \frac{323 \text{ K}}{298 \text{ K}} = 163 \text{ mL} \rightarrow 160 \text{ mL}$$

2. The pressure in a bicycle tire is 105 psi at 25°C here in Fresno. You take the bicycle up to Huntington, where the temperature is –5°C. What is the pressure in the tire? (V and n = constant)

	P	T	V	n
Initial	105psi	298K	--	--
Final	?	268K	--	--
Effect	↓	↓	--	--

$$P_2 = 105 \text{ psi} \times \frac{268 \text{ K}}{298 \text{ K}} = 94.4 \text{ psi}$$

3. What would be the new pressure if 250 cm<sup>3</sup> of gas at standard pressure is compressed to a volume of 150 cm<sup>3</sup>? (T, n = constant)

	P	T	V	n
Initial	1atm	--	250cm <sup>3</sup>	--
Final	?	--	150cm <sup>3</sup>	--
Effect	↑	--	↓	--

$$P_2 = 1 \text{ atm} \times \frac{250 \text{ cm}^3}{150 \text{ cm}^3} = 1.7 \text{ atm}$$

or 1300 mmHg  
or 170 kPa  
or 25 psi

# Chemistry Unit 2 Pvtn Problems Answers

**Ian Pickup**

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