

# Moles and Reactions

Student

Date

Page

## Introduction: Molar Ratios in Equations

Proportive quantities may be obtained from the ratios shown in the balanced equation for the combination of proper units.



The mole number of product that obtained, when units, such as moles, are substituted for the coefficients of reactants or products, gives the mole number of product that is obtained. Similarly, the mole number of product that reacts when units are substituted for the coefficients of reactants or products, gives the mole number of product that reacts when units are substituted for the coefficients of reactants or products. The number of moles of product that is obtained, when units are substituted for the coefficients of reactants or products, gives the mole number of product that is obtained.

## Critical Thinking Questions

Using the equations 1-4, either write balanced equations for the combination of proper units.

1. a) How many moles of water are produced when 1.00 moles of hydrogen are reacted?



- b) How many moles of water are produced when 1.00 moles of oxygen are reacted?



2. a) How many moles of water are produced when 1.00 moles of hydrogen are reacted?



3. a) How many moles of water are produced when 1.00 moles of oxygen are reacted?

The balanced equation represents the reaction of 2 moles of hydrogen and 1 mole of oxygen to produce 2 moles of water.

4. a) How many moles of water are produced when 1.00 moles of hydrogen are reacted?



- b) How many moles of water are produced when 1.00 moles of oxygen are reacted?



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