

Covalent Bonding

How does a covalent bond form?

- A **covalent bond** forms when two atoms **share** one or more pairs of electrons.
- Atoms in covalent bonds typically have **similar electronegativities** and are **usually non-metals**.
- The shared electrons allow each atom to achieve a full outer electron shell, similar to the noble gas configuration.

Types of Covalent Bonds

- **Single Bond:** One pair of electrons is shared (e.g., H_2 , Cl_2).
- **Double Bond:** Two pairs of electrons are shared (e.g., O_2 , CO_2).
- **Triple Bond:** Three pairs of electrons are shared (e.g., N_2).



Dot-and-Cross Diagrams

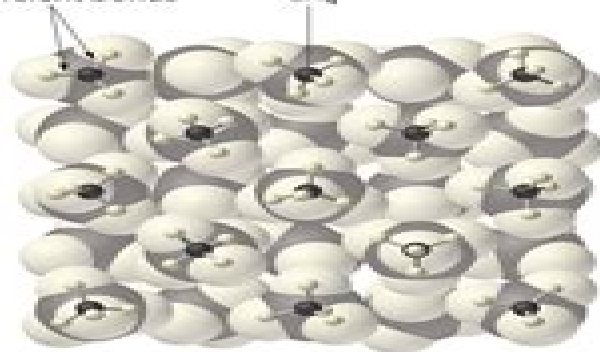
- Only the outer electron shells (valence electrons) are shown.
- Shared pairs of electrons between atoms are shown as pairs of dots and crosses.
- It represents how the atoms bond and achieve stable electron configurations.

Example: Water H_2O



Covalent bonds

CH_4



Properties of Covalent Compounds

- **Molecular Structure:** They are made up of discrete molecules with **weak intermolecular forces**.
- **Melting and Boiling Points:** Generally **lower** than ionic compounds because the forces between molecules are weaker.
- **Electrical Conductivity:** **Poor conductors** of electricity in all states, as they do not have free ions or electrons.
- **Solubility:** Often **soluble in nonpolar solvents** and may be insoluble in water, depending on their polarity.

Covalent Bonding Exploration Guide Answers

E Durkheim



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Computational Approaches for Chemistry Under Extreme Conditions Nir Goldman, 2019-02-18 This book presents recently developed computational approaches for the study of reactive materials under extreme physical and thermodynamic conditions. It delves into cutting edge developments in simulation methods for reactive materials including quantum calculations spanning nanometer length scales and picosecond timescales to reactive force fields, coarse grained approaches and machine learning methods spanning microns and nanoseconds and beyond. These methods are discussed in the context of a broad range of fields including prebiotic chemistry in impacting comets, studies of planetary interiors, high pressure synthesis of new compounds and detonations of energetic materials. The book presents a pedagogical approach for these state of the art approaches compiled into a single source for the first time. Ultimately the volume aims to make valuable research tools accessible to experimentalists and theoreticians alike for any number of scientific efforts spanning many different types of compounds and reactive conditions. *Fullerenes and Other Carbon-Rich Nanostructures* Jean-François Nierengarten, 2014-05-16 Yanfei Shen and Takashi Nakanishi Exotic Self Organized Fullerene Materials Based on Uncommon Hydrophobic Amphiphilic Approach Yuming Zhao and Guang Chen C60 Fullerene Amphiphiles as Supramolecular Building Blocks for Organized and Well Defined Nano scale Objects Anna Troeger Vito Sgobba and Dirk M Guldi Multilayer Assembly for Solar Energy Conversion Delphine Felder Flesch Self or Induced Organization of 60 Fullerene Hexakisadducts Andr s de la Escosura Olga Trukhina and Tom s Torres Dual Role of Phthalocyanines in Carbon Nano structure Based Organic Photovoltaics Riccardo Marega Davide Giust and Davide Bonifazi Supramolecular Chemistry of Carbon Nano tubes at Interfaces Toward Applications Stephanie Frankenberger Johanna A Januszewski and Rik R Tykwinski Oligomers from sp Hybridized Carbon Cumulenes and Polyynes

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the platform or they can modify the questions or author them from scratch Not only does the software allow students to easily draw and interact with structures it allows them to draw entire reaction mechanisms including showing the movement of electrons with curved electron arrows

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