



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Draw a molecular formula for each of the following compounds.

- a. Silicon tetrachloride b. Nitrogen dioxide c. Hydrogen peroxide
d. Carbon disulfide e. Sulfur dioxide

2. Write the formula for the following molecular compounds.

- a. Carbon tetrachloride b. Nitrogen dioxide c. Hydrogen peroxide
d. Carbon disulfide e. Sulfur dioxide

3. Draw a Lewis structure for each of the following molecular compounds.

a. Hydrogen chloride b. Carbon dioxide c. Nitrogen gas d. Water e. Ammonia f. Hydrogen cyanide g. Hydrogen sulfide h. Hydrogen bromide i. Hydrogen iodide

4. Draw a Lewis structure for each of the following molecular compounds. Label each atom as being an electrophile or a nucleophile.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw a Lewis structure for hydrogen fluoride, HF.



2. Draw a Lewis structure for methane, CH₄.



3. Draw a Lewis structure for carbon dioxide, CO₂.



4. Draw a Lewis structure for acetylene, C₂H₂.

5. Draw a Lewis structure for each of the following compounds. Label each atom as being an electrophile or a nucleophile.



6. Draw a Lewis structure for each of the following compounds.

SECTION 8.3 ELECTRONIC STRUCTURE

1. Write the electron configuration for each of the following elements: a. H b. He c. Li d. Be e. B f. C g. N h. O i. F j. Ne

2. Write the electron configuration for each of the following elements: a. Na b. Mg c. Al d. Si e. P f. S g. Cl h. Ar

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