

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

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Karen C. Timberlake



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Table of Contents Covalent Compound Study Guide Answers

1. Understanding the eBook Covalent Compound Study Guide Answers
 - The Rise of Digital Reading Covalent Compound Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Compound Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Compound Study Guide Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Covalent Compound Study Guide Answers
 - Personalized Recommendations
 - Covalent Compound Study Guide Answers User Reviews and Ratings
 - Covalent Compound Study Guide Answers and Bestseller Lists
5. Accessing Covalent Compound Study Guide Answers Free and Paid eBooks
 - Covalent Compound Study Guide Answers Public Domain eBooks
 - Covalent Compound Study Guide Answers eBook Subscription Services
 - Covalent Compound Study Guide Answers Budget-Friendly Options
6. Navigating Covalent Compound Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Compound Study Guide Answers Compatibility with Devices
 - Covalent Compound Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Compound Study Guide Answers
 - Highlighting and Note-Taking Covalent Compound Study Guide Answers
 - Interactive Elements Covalent Compound Study Guide Answers
8. Staying Engaged with Covalent Compound Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Compound Study Guide Answers
9. Balancing eBooks and Physical Books Covalent Compound Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Compound Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Covalent Compound Study Guide Answers
 - Setting Reading Goals Covalent Compound Study Guide Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Covalent Compound Study Guide Answers
 - Fact-Checking eBook Content of Covalent Compound Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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