

# GENERAL CHEMISTRY REFERENCE SHEET

| Common Metric Prefixes and Units |        |                                             |
|----------------------------------|--------|---------------------------------------------|
| Unit                             | Symbol | Conversions                                 |
| meter                            | m      | 1 m = 39.37 inches                          |
| centimeter                       | cm     | 1 inch = 2.54 cm<br>exactly                 |
| Angstrom                         | Å      | 1 Å = $1 \times 10^{-10}$ m                 |
| gram                             | g      | 1 g = 0.002205 lb                           |
| liter                            | L      | 1 L = 1000 cm <sup>3</sup><br>= 1.056710 qt |
| Kelvin                           | K      | K = °C + 273.15                             |
| kilo                             | k      | 10 <sup>3</sup>                             |
| deci                             | d      | 10 <sup>-1</sup>                            |
| centi                            | c      | 10 <sup>-2</sup>                            |
| milli                            | m      | 10 <sup>-3</sup>                            |
| micro                            | μ      | 10 <sup>-6</sup>                            |
| nano                             | n      | 10 <sup>-9</sup>                            |
| pico                             | p      | 10 <sup>-12</sup>                           |

## Nomenclature

| Common Cations |                                    |                          |
|----------------|------------------------------------|--------------------------|
| Charge         | Formula                            | Name                     |
| 1+             | H <sup>+</sup>                     | hydrogen ion             |
|                | H <sub>3</sub> O <sup>+</sup>      | hydronium ion            |
|                | Alkali metal <sup>+</sup>          | alkali metal ion         |
|                | NH <sub>4</sub> <sup>+</sup>       | ammonium ion             |
|                | Ag <sup>+</sup>                    | silver ion               |
| 2+             | Cu <sup>+</sup>                    | copper(I)                |
|                | Alkaline earth metal <sup>2+</sup> | alkaline earth metal ion |
|                | Zn <sup>2+</sup>                   | zinc ion                 |
|                | Co <sup>2+</sup>                   | cobalt(II)               |
|                | Cu <sup>2+</sup>                   | copper(II)               |
|                | Fe <sup>2+</sup>                   | iron(II)                 |
|                | Mn <sup>2+</sup>                   | manganese(II)            |
|                | Hg <sub>2</sub> <sup>2+</sup>      | mercury(I)               |
|                | Hg <sub>2</sub> <sup>2+</sup>      | mercury(II)              |
|                | Ni <sup>2+</sup>                   | nickel(II)               |
|                | Pb <sup>2+</sup>                   | lead(II)                 |
|                | Sn <sup>2+</sup>                   | tin(II)                  |
| 3+             | Al <sup>3+</sup>                   | aluminum ion             |
|                | Cr <sup>3+</sup>                   | chromium(III)            |
|                | Fe <sup>3+</sup>                   | iron(III)                |

| Common Anions |                               |                                       |
|---------------|-------------------------------|---------------------------------------|
| Charge        | Formula                       | Name                                  |
| 1-            | H <sup>-</sup>                | hydride ion                           |
|               | F <sup>-</sup>                | fluoride ion                          |
|               | Cl <sup>-</sup>               | chloride ion                          |
|               | Br <sup>-</sup>               | bromide ion                           |
|               | I <sup>-</sup>                | iodide ion                            |
|               | CN <sup>-</sup>               | cyanide ion                           |
|               | OH <sup>-</sup>               | hydroxide ion                         |
|               | NO <sub>3</sub> <sup>-</sup>  | nitrate                               |
|               | NO <sub>2</sub> <sup>-</sup>  | nitrite                               |
|               | HCO <sub>3</sub> <sup>-</sup> | hydrogen carbonate or bicarbonate ion |
|               | HSO <sub>4</sub> <sup>-</sup> | hydrogen sulfate ion                  |

| Charge | Formula                                                                                             | Name                     |
|--------|-----------------------------------------------------------------------------------------------------|--------------------------|
| 1-     | ClO <sub>4</sub> <sup>-</sup>                                                                       | perchlorate ion          |
|        | ClO <sub>3</sub> <sup>-</sup>                                                                       | chlorate ion             |
|        | ClO <sub>2</sub> <sup>-</sup>                                                                       | chlorite ion             |
|        | ClO <sup>-</sup>                                                                                    | hypochlorite ion         |
|        | MnO <sub>4</sub> <sup>-</sup>                                                                       | permanganate ion         |
|        | H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>                                                         | dihydrogen phosphate ion |
|        | C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup><br>(or CH <sub>3</sub> COO <sup>-</sup> ) | acetate ion              |
| 2-     | O <sup>2-</sup>                                                                                     | oxide ion                |
|        | S <sup>2-</sup>                                                                                     | sulfide ion              |
|        | SO <sub>4</sub> <sup>2-</sup>                                                                       | sulfate ion              |
|        | SO <sub>3</sub> <sup>2-</sup>                                                                       | sulfite ion              |
|        | CO <sub>3</sub> <sup>2-</sup>                                                                       | carbonate ion            |
|        | Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>                                                        | dichromate ion           |
| 3-     | HPO <sub>4</sub> <sup>2-</sup>                                                                      | hydrogen phosphate       |
|        | PO <sub>4</sub> <sup>3-</sup>                                                                       | phosphate ion            |

## Other Compounds

| Formula                                                                                                                   | Name                  |
|---------------------------------------------------------------------------------------------------------------------------|-----------------------|
| H <sub>2</sub> O                                                                                                          | water                 |
| NH <sub>3</sub>                                                                                                           | ammonia               |
| CH <sub>4</sub>                                                                                                           | methane (natural gas) |
| O <sub>3</sub>                                                                                                            | ozone                 |
| H <sub>2</sub> O <sub>2</sub>                                                                                             | hydrogen peroxide     |
| H <sub>2</sub> , N <sub>2</sub> , O <sub>2</sub> , F <sub>2</sub> , Cl <sub>2</sub> ,<br>Br <sub>2</sub> , I <sub>2</sub> | diatomic molecules    |
| He, Ne, Ar                                                                                                                | Noble gases           |

## Acids

| Formula                                                                    | Name                       |
|----------------------------------------------------------------------------|----------------------------|
| HNO <sub>3</sub>                                                           | nitric acid (strong)       |
| H <sub>2</sub> SO <sub>4</sub>                                             | sulfuric acid (strong)     |
| HCl                                                                        | hydrochloric acid (strong) |
| HBr                                                                        | hydrobromic acid (strong)  |
| HI                                                                         | hydroiodic acid (strong)   |
| HClO <sub>4</sub>                                                          | perchloric acid (strong)   |
| H <sub>3</sub> PO <sub>4</sub>                                             | phosphoric acid (weak)     |
| H <sub>2</sub> CO <sub>3</sub>                                             | carbonic acid (weak)       |
| HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> (or<br>CH <sub>3</sub> COOH) | acetic acid (weak)         |
| HF                                                                         | hydrofluoric acid          |

## Bases

| Formula             | Name                         |
|---------------------|------------------------------|
| NaOH                | sodium hydroxide (strong)    |
| KOH                 | potassium hydroxide (strong) |
| Mg(OH) <sub>2</sub> | magnesium hydroxide (strong) |
| Ca(OH) <sub>2</sub> | calcium hydroxide (strong)   |
| NH <sub>3</sub>     | ammonia (weak)               |
| NR <sub>3</sub>     | other amines (weak)          |

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*Book of ASTM Standards, with Related Material* American Society for Testing and Materials, 1976      **Annual Book of ASTM Standards** American Society for Testing and Materials, 1978 Index to ASTM standards issued as last part of each vol

**EPA-600/2**, 1979      **Research Reporting Series**, 1979      Tables of Molecular Vibrational Frequencies Takehiko Shimanouchi, 1972      Evaluation of Dewatering Devices for Producing High-solids Sludge Cake Alan F. Cassel, Berinda P. Johnson, 1979      *Next Generation Kinase Inhibitors* Paul Shapiro, 2020-07-14 Protein kinases are fascinating enzymes that maintain the proper function of nearly every task performed by the cells of the human body By extracting a phosphate from the energy molecule ATP and linking it to another protein protein kinases alter the structure and ultimate function of other proteins In this way protein kinases help monitor the extracellular environment and integrate signaling cues that for the most part are beneficial for human health and survival However protein kinases are often dysregulated and responsible for the initiation and progression of many types of cancers inflammatory disorders and other diseases Thus decades of research have revealed much about how protein kinases are regulated and approaches to inhibit these enzymes to treat disease However nearly 30 years since the identification of the first clinically beneficial small molecule protein kinase inhibitor there are only a few examples where these drugs provide sustained and durable patient responses The goal of this book is to provide biomedical scientists graduate and professional degree students insight into different approaches using small molecules to block specific protein kinase functions that promote disease      *Reader's Index and Guide*, 1917      **The Reader's Index & Guide** Croyden Public Libraries, 1930      **Reference Catalogue of Current Literature**, 1902      **The New Britannica/Webster Dictionary & Reference Guide** Encyclopaedia Britannica, Inc, 1981      The Library Catalogs of the Hoover Institution on War, Revolution, and Peace, Stanford University Hoover Institution on War, Revolution, and Peace, 1969      **National Union Catalog**, 1980 Includes entries for maps and atlases      *The National Union Catalog, Pre-1956 Imprints*, 1974      **Information Resources in Toxicology** P.J. Bert Hakkinen, Asish Mohapatra, Steven G. G. Gilbert, 2009-08-19 This latest version of Information Resources in Toxicology IRT continues a tradition established in 1982 with the publication of the first edition in presenting an extensive itemization review and commentary on the information infrastructure of the field This book is a unique wide ranging international annotated bibliography and compendium of major resources in toxicology and allied fields such as environmental and occupational health chemical safety and risk assessment Thoroughly updated the current edition analyzes technological changes and is rife with online tools and links to Web sites IRT IV is highly structured providing easy access to its information Among the hot topics covered are Disaster Preparedness and Management Nanotechnology Omics the Precautionary Principle Risk Assessment and Biological Chemical and Radioactive Terrorism and Warfare are among the designated International in scope with contributions from over 30 countries Numerous key references and relevant Web links Concise narratives about toxicologic sub disciplines Valuable

appendices such as the IUPAC Glossary of Terms in Toxicology Authored by experts in their respective sub disciplines within toxicology    *NBS Special Publication* ,1968    **Monthly Catalog of United States Government Publications** ,    The Chemical News ,1917    **The Reader's Index** Croydon Public Libraries,1917    The National Union Catalogs, 1963- ,1964

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240 pixel density 167 dual screen support no scra

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