

Physical Science – Chapter 18 Study Guide Answers

Chapter Objectives

1. Define pure substances and mixtures.

Pure substances – type of matter with a fixed composition, element or compound (composition is definite)

Mixtures – 2 or more substances that can be separated by physical means (Composition is variable)

2. Identify elements and compounds.

Elements – pure substances in which all atoms have the same identity

Compounds – when 2 or more different elements combine in a fixed proportion

3. Compare and contrast solutions, colloids, and suspensions.

Solutions – particles are so small they cannot be seen with a microscope and will never settle out; light cannot pass through, homogenous mixtures

Colloids – particles that are larger than those in solution but not heavy enough to settle out; Ex: milk, light passes through/scatters (Tyndall effect)

Suspensions – heterogeneous mixtures, contains liquid, particles will settle if left undisturbed, Ex: pond water; will scatter light

4. Describe the phase and energy changes associated with STATE changes.

Changing states of matter (from solids, liquids, or gas to another state), physical changes

Melting – solid to a liquid, add energy

Freezing: liquid to solid, releases energy

Vaporization: liquid to gas, add energy

- Boiling: throughout the liquid

- Evaporation: surface only

Condensation: gas to liquid, releases energy

Sublimation: solid to gas, adds energy

5. Identify substances using physical properties.

Physical properties – properties than can be observed without altering identity of substance

Alterations in form but not identity

Density, freezing/melting point, color

6. Compare and contrast physical and chemical changes.

Ch 18 Study Guide Answers

John R. Schermerhorn



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