

Chapter 18

Study Guide - Answers

1. Write the expression for the equilibrium constant for each of the following reactions.



$$K_{\text{eq}} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$



$$K_{\text{eq}} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$



$$K_{\text{eq}} = \frac{[\text{H}^+][\text{CN}^-]}{[\text{HCN}]}$$



$$K_{\text{eq}} = \frac{[\text{H}^+][\text{HSO}_4^-]}{[\text{H}_2\text{SO}_4]}$$



$$K_{\text{eq}} = [\text{Pb}^{2+}][\text{F}^-]^2$$

2. From the data provided below, calculate the value of the equilibrium constant for the reaction.



$$[\text{H}_2] = [\text{Cl}_2] = 1.0 \times 10^{-2}; [\text{HCl}] = 1.0 \times 10^{-4}$$

$$K_{\text{eq}} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$

$$K_{\text{eq}} = \frac{[1.0 \times 10^{-4}]^2}{[1.0 \times 10^{-2}][1.0 \times 10^{-2}]} = 0.0001$$



$$[\text{N}_2] = 4.4 \times 10^{-2}; [\text{H}_2] = 1.2 \times 10^{-1}; [\text{NH}_3] = 3.4 \times 10^{-3}$$

$$K_{\text{eq}} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$

$$K_{\text{eq}} = \frac{[3.4 \times 10^{-3}]^2}{[4.4 \times 10^{-2}][1.2 \times 10^{-1}]^3} = 0.152$$

3. For the reaction $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$, $K_{\text{eq}} = 0.2$. At a particular time, the following concentrations are measured: $[\text{N}_2\text{O}_4] = 2.0 \text{ M}$, $[\text{NO}_2] = 0.2 \text{ M}$. Is this reaction at equilibrium? If not, in which direction will the reaction proceed?



$$[\text{N}_2\text{O}_4] = 2.0 \text{ M}$$

$$[\text{NO}_2] = 0.2 \text{ M}$$

$$Q = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$

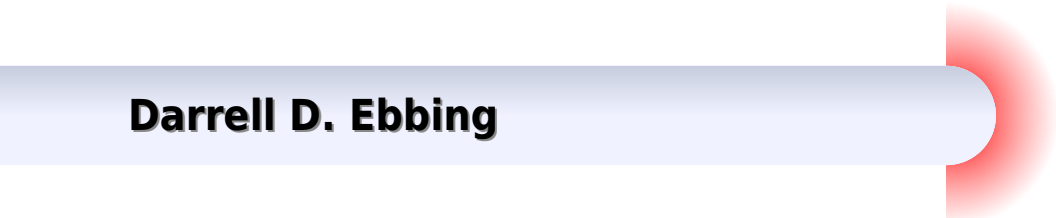
$$Q = \frac{[0.2]^2}{[2.0]} = 0.02$$

$$K_{\text{eq}} = 0.2$$

$Q < K_{\text{eq}}$, so the reaction will proceed to the right (forward).

Ch 18 Study Guide Answers For Chemistry

Darrell D. Ebbing



Ch 18 Study Guide Answers For Chemistry:

Study Guide with Answers to Selected Problems H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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Study Guide, Answer Book to Accompany Organic Chemistry Alan S. Wingrove, Robert L. Caret, 1981

Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

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Organic Chemistry, Study Guide and Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, 2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and

breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science *Study Guide for Chemical Principles [by] Steven S. Zumdahl* Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples *A Study Guide to Basic Principles of Organic Chemistry* John D. Roberts, Marjorie C. Caserio, 1977 *Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual* T. W. Graham Solomons, 1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

Instructor's Manual and Study Guide Answers for the Human Body in Health and Disease Barbara Janson Cohen, Memmler, 1996 *Chemistry, Student Study Guide* John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that s real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You ll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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