

Concentrations and Dilutions: How to Dilute a Concentrated Solution

In the first part of this lab, you learned about the actions of solutes added to water to form solutions of various concentrations. In this exercise, you will use a powerful and simple tool to determine the concentration of a solution when it has been diluted with additional water or made more concentrated due to evaporation of water.

Some basic equations for dilution effects:

Molarity:

Dilution:

Evaporation:

Procedure: Print this page with the other AP Chemistry AP Concentrations [Worksheet](#)

Point 1: Dilution

At this point you should be familiar with the Concentration simulation. Because some items may have different labels or have mistakes in the original simulation, pay particular attention to the effect of evaporation and addition of water to the solution's concentration.



How does the concentration change as additional water is added?

Why?



How does the concentration change as evaporation occurs?

Why?

Does evaporation change the concentration of a concentrated solution?

Why is this?

Why is this?



Using the concentrated solution again, add 10 L of Distilled Water to an empty beaker. What is the concentration?

Is this solution saturated?

How do you know?

Fill the beaker with another 10 L of water. What is the new concentration?

Complete the table below, using **Concentrated** as an empty beaker, writing the concentrated solution in the boxes provided.

Dilute 0.25L of 25% solution	25% solution + 0.25L water	0.25L solution + 0.25L water	25% solution + 0.25L water

Repeat the exercise, using **Concentrated** as an empty beaker.

Dilute 0.25L of 25% solution	25% solution + 0.25L water	0.25L solution + 0.25L water	25% solution + 0.25L water

What do you notice about the concentration change as each addition of 0.25L of water is added to the concentrated 25% solution?



Concentrated water

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Vu, Phu,Fredrickson, Scott,Moore, Carl

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Table of Contents Concentration And Molarity Phet Chemistry Labs Key

1. Understanding the eBook Concentration And Molarity Phet Chemistry Labs Key
 - The Rise of Digital Reading Concentration And Molarity Phet Chemistry Labs Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Concentration And Molarity Phet Chemistry Labs Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Concentration And Molarity Phet Chemistry Labs Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Concentration And Molarity Phet Chemistry Labs Key
 - Personalized Recommendations
 - Concentration And Molarity Phet Chemistry Labs Key User Reviews and Ratings
 - Concentration And Molarity Phet Chemistry Labs Key and Bestseller Lists
5. Accessing Concentration And Molarity Phet Chemistry Labs Key Free and Paid eBooks
 - Concentration And Molarity Phet Chemistry Labs Key Public Domain eBooks
 - Concentration And Molarity Phet Chemistry Labs Key eBook Subscription Services
 - Concentration And Molarity Phet Chemistry Labs Key Budget-Friendly Options
6. Navigating Concentration And Molarity Phet Chemistry Labs Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Concentration And Molarity Phet Chemistry Labs Key Compatibility with Devices
 - Concentration And Molarity Phet Chemistry Labs Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concentration And Molarity Phet Chemistry Labs Key
 - Highlighting and Note-Taking Concentration And Molarity Phet Chemistry Labs Key
 - Interactive Elements Concentration And Molarity Phet Chemistry Labs Key
8. Staying Engaged with Concentration And Molarity Phet Chemistry Labs Key

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concentration And Molarity Phet Chemistry Labs Key
9. Balancing eBooks and Physical Books Concentration And Molarity Phet Chemistry Labs Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concentration And Molarity Phet Chemistry Labs Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Concentration And Molarity Phet Chemistry Labs Key
- Setting Reading Goals Concentration And Molarity Phet Chemistry Labs Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Concentration And Molarity Phet Chemistry Labs Key
- Fact-Checking eBook Content of Concentration And Molarity Phet Chemistry Labs Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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