

Conceptual Chemistry

PRACTICE PAGE

Chapter 10: Acids and Bases In Our Environment

Expert-Only Questions ♦♦♦

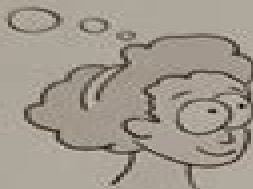
1. To say that an acid is strong in aqueous solution means that the acid
- produces a large concentration of hydronium ions.
 - is quite odorous when dissolved in water.
 - does not form chemical bonds with water.
 - does not separate into ions when dissolved in water.

$$10^3 = 1000$$

$$10^2 = 100$$

$$10^1 = 10$$

$$10^0 = ?$$



2. What does it mean to say that a solution has a pH of zero?
- There are no hydronium ions in that solution.
 - The concentration of hydronium ions is 1 M.
 - The solution is very alkaline.
 - It is not possible to have a pH equal to zero!

3. We measure the "acidity" of a solution by measuring the concentration of hydronium ions, H_3O^+ . What does it mean to say that one has a "neutral" solution, such as one has with pure water?

- There are no hydronium ions in this solution.
- The concentration of hydronium ions is very small.
- The concentration of hydronium ions equals the concentration of hydroxide ions, OH^- .



Pure?



4. Water reacting with itself is an endothermic reaction:



What happens, therefore, to the pH of neutral water as its temperature is increased and thus more thermal energy becomes available? The pH of neutral water

- remains at 7 because the pH of neutral water is always 7.
- decreases to some value below 7.
- increases to some value above 7.

The reaction of water with itself is favored by greater temperatures and pressures. At the super-high temperature and pressure of 374°C and 218 atmospheres, the reaction of water with water is so great that K_w equals 10^{-13} , which is ten times greater than the usual value of 10^{-14} . What is the pH of such a solution? Is this pH acidic, neutral, or basic? What might you predict about the corrosive properties of water under these conditions? So much to learn about our universe!



S. Subak

Conceptual Chemistry Practice Page Answers

Dr. Rastogi, Goyal



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Conceptual Physical Science Paul G. Hewitt, 1998-11 Conceptual Physical Science Third Edition takes learning physical science to a new level by combining Hewitt's leading conceptual approach and friendly writing style in a new edition that provides stronger integration of the sciences more quantitative coverage and a wealth of new media resources to help readers. The dynamic new media program includes hundreds of animations and interactive tutorials developed specifically for students taking physical science courses. Media references throughout the book point readers to additional online help. **KEY TOPICS** The book's consistent high quality coverage includes five new chapters on chemistry, astronomy, and earth science for an even more balanced approach to physical science. For college instructors, students, or anyone interested in physical science.

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Basic Chemistry Steven S. Zumdahl, 2004 Description Not Yet Available **The Chemistry of Everything** Kimberley Waldron, 2007 The Chemistry of Everything addresses the need to know basics of chemistry required to grasp everyday science issues Through innovative themes and creative applications it provides an engaging introduction to chemistry for nonscience majors Mixes basic chemical principles from physical inorganic organic analytical and biological specializations to support thematic coverage of topics such as diamonds groceries and drugs Extends readers vocabulary and knowledge of the scientific issues encountered in daily life Addresses issues of ethics and responsible use in contemporary science Captures the current fascination with forensics through Chemistry at the Crime Scene boxed sections For those interested in basic chemistry **Publisher's Monthly**, 1998

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