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protection inappropriate use of emotions invasion of privacy autonomy suppression decrease in human interaction and cognitive safety Given the novelty of these technologies and the uncertainties surrounding the impact of care automation it is unclear how the law should respond This book investigates the legal and regulatory implications of the growing use of personal care robots for healthcare purposes It explores the interplay between various aspects of the law including safety data protection responsibility transparency autonomy and dignity and it examines different robotic and AI systems such as social therapy robots physical assistant robots for rehabilitation and wheeled passenger carriers Highlighting specific problems and challenges in regulating complex cyber physical systems in concrete healthcare applications it critically assesses the adequacy of current industry standards and emerging regulatory initiatives for robots and AI After analyzing the potential legal and ethical issues associated with personal care robots it concludes that the primarily principle based approach of recent law and robotics studies is too abstract to be as effective as required by the personal care context Instead it recommends bridging the gap between general legal principles and their applicability in concrete robotic and AI technologies with a risk based approach using impact assessments As the first book to compile both legal and regulatory aspects of personal care robots this book will be a valuable addition to the literature on robotics artificial intelligence human robot interaction law and philosophy of technology

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web jan 30 2020 ans the molality of solution is 1 mol kg 1 or 1 m example 02 11 11 g of urea  $nh_2CONH_2$  was dissolved in 100 g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11 11 g mass of solvent water 100 g 0 1 kg to find molarity of solution solution

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0.449 molal  $KBr$  that contains 2.92 kg of water 1 molality moles solute divided by kilograms solute 0.449 mol kg x 2.92 kg x 1.31108 mol of  $KBr$  2 moles times molar mass equals grams 1.31108 mol 119.0023 g mol 156 g  $KBr$

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of

caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web a derivatization that achieves this goes as molarity moles of solute volume of solution in liter moles molarity  $\times$  volume in liter mass moles  $\times$  molar mass numerical answers 1

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web solution a 3.8  $\times$  10.51 grams b 4.11  $\times$  10.45 grams c 7.31  $\times$  10.25 grams 4 7 1 mole practice with guidance exercises is shared under a not declared license and was authored remixed and or curated by libretexts these are homework exercises to accompany the textmap created for chemistry a molecular approach by nivaldo tro

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web what is the mole fraction of licl in the solution 11.8 practice problem acetone c 3 h 6 o has a density of 0.791 g/ml and xylene c 8 h 10 has a density of 0.864 g/ml a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if 19.0 ml of acetone was dissolved in 98.4 ml of xylene 7.1

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web since the mole fraction of the o<sub>2</sub> is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.21 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of o<sub>2</sub> are mixed in a flask at room temperature

