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**David Engle,Cisco Learning Institute**

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web what is an oxidation reduction or redox reaction this is a reaction in which electrons are transferred between reactants what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 fe2 co co2 fe oxidation co co2 2e reduction fe2 2e fe 24 3 ag ni ni3 3 ag oxidation reduction 25 cu2 pb pb2 cu oxidation reduction 26 o 2

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a i2o5 5co i2 5co2 i 2 o 5 5 c o i 2 5 c o 2

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also 2sr o2 2sro sr0 to sr2 oxidized reducing agent o0 to o2 reduced ox ag 2li s li2s li0 to li1 oxidized red ag s0 to s2 reduced ox ag 3 2cs br2