

Calorimetry Practice Problems

1. How much energy is needed to change the temperature of 50.0 g of water by 15.0°C?
2. How many grams of water can be heated from 20.0 °C to 75°C using 12500.0 Joules?
3. What is the final temperature after 840 Joules is absorbed by 10.0g of water at 25.0°C?
4. The heat capacity of aluminum is 0.900 J/g°C.
 - a. How much energy is needed to raise the temperature of a 8.50×10^2 g block of aluminum from 22.8°C to 94.6°C?
5. A piece of metal weighing 59.047 g was heated to 100.0 °C and then put it into 100.0 mL of water (initially at 23.7 °C). The metal and water were allowed to come to an equilibrium temperature, determined to be 27.8 °C. Assuming no heat lost to the environment, calculate the specific heat of the metal. (Hint: First calculate the heat absorbed by the water then use this value for "Q" to determine the specific heat of the metal in a second calculation)
6. In a coffee-cup calorimeter, 100.0 g of H₂O and 100.0 mL of HCl are mixed. The HCl had an initial temperature of 44.6 °C and the water was originally at 24.6 °C. After the reaction, the temperature of both substances is 31.3 °C.
 - a. Was the reaction exothermic or endothermic? Explain.
 - b. Calculate how much heat the water lost or gained.

Calorimetry Problems With Answers

**Patrick E. McMahon, Rosemary
McMahon, Bohdan Khomtchouk**



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

A Liquid-tin Solution Calorimeter and Its Application to the Problem of Bonding Energies in Alloys Alfred Goldberg, 1955

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