



balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay

electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- Sulfur hexafluoride SF_6
- Lithium nitride Li_3N
- Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- Tin (II) chloride SnCl_2
- Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- Mercury (I) chloride Hg_2Cl_2
- Potassium bromate KBrO_3
- Hydrosulfuric acid $\text{H}_2\text{S}(\text{aq})$
- Chloric acid HClO_3
- Sulfurous acid H_2SO_3

Name each of the following compounds.

- CuSO_4 Copper (II) sulfate
- AlF_3 aluminum fluoride
- HI hydrogen iodide
- NO nitrogen monoxide
- H_2Se hydroseonic acid
- HNO_3 nitric acid
- HNO_2 nitrous acid
- NaHSO_4 sodium bisulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25 \quad \boxed{61.5 \text{ yr}}$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} \boxed{12.5 \text{ mg}}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 44 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

Chemistry Study Guide With Answers

**Denise Guinn, Rebecca Brewer, Rachel
C. Lum**

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Organic Chemistry, Study Guide and Solutions Manual T. W. Graham Solomons,Craig B. Fryhle,2000

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Chemistry Steven S. Zumdahl,2000

[Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice](#) Barron's Educational Series,Mark Kernion,Joseph A. Mascetta,2021-09-07 Barron s Science 360 Chemistry is your complete go to guide for everything chemistry This comprehensive guide is an essential resource for High school and college courses Homeschooling Virtual Learning Learning pods Inside you ll find Comprehensive Content Review Begin your study with the basic building block of chemistry and build as you go Topics include atomic structure chemical formulas electrochemistry the basics of organic chemistry and much more Effective Organization Topic organization and simple lesson formats break down the subject

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[Student Study Guide for Chemistry](#) John E. McMurry, Robert C. Fay, Jill Kirsten Robinson, Julie Klare, Donnajean Fredeen, James Zubricky, 2015-03-09

The Study Guide includes learning goals, an overview, a review section with worked examples and self tests with answers.

Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan

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Introduction to Chemistry Mark Bishop, 2001-12

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