

Concentration Worksheet – Answers

- 1) How many grams of beryllium chloride are needed to make 125 mL of a 0.050 *M* solution?
0.50 grams
- 2) How many grams of beryllium chloride would you need to add to 125 mL of water to make a 0.050 molal solution?
0.50 grams
- 3) The density of ethanol is 0.789 g/mL. How many grams of ethanol should be mixed with 225 mL of water to make a 4.5% (v/v) mixture?
7.99 grams
- 4) Explain how to make at least one liter of a 1.25 molal ammonium hydroxide solution.
Dissolve 43.8 grams (1.25 moles) of ammonium hydroxide in 1 L H₂O.
- 5) What is the molarity of a solution in which 0.45 grams of sodium nitrate are dissolved in 265 mL of solution.
0.020 *M*
- 6) What is the mole fraction of sulfuric acid in a solution made by adding 3.4 grams of sulfuric acid to 3,500 mL of water?
 1.8×10^{-4}
- 7) What will the volume of a 0.50 *M* solution be if it contains 25 grams of calcium hydroxide?
680 mL
- 8) How many grams of ammonia are present in 5.0 L of a 0.050 *M* solution?
4.3 grams

Chemistry Study Guide Solution Concentration Answers

Bernice G. Segal, Peter S. Shenkin



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Organic Chemistry Study Guide and Solutions Marc Loudon, Jim Parise, 2015-07-01 Parise and Loudon's Study Guide and Solutions Manual offers the following learning aids Links that provide hints for study approaches to problem solving and additional explanations of challenging topics Further Explorations that provide additional depth on key topics Reaction summaries that delve into key mechanisms and stereochemistry Solutions to all the textbook problems Rather than providing just the answer many of the solutions provide detailed explanations of how the problem should be approached *Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl* Paul B. Kelter, 1993 **Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application **Introduction to Chemistry, Study Guide** T. R. Dickson, 1999-07-07 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and chemical vision and refining problem solving skills Many chapters in this edition has been rewritten and rearranged to vitalize the topics and to include interesting examples analogies and images **Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry** Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface *Chemistry, Student Study Guide* John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased

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