

## Information: Orbitals

So far we have learned that inside energy levels there are different sublevels. Now we will look at orbitals. Orbitals are located inside sublevels just like streets are located inside subdivisions. Different sublevels have different numbers of orbitals.

TABLE 2

Sublevel	# of Orbitals Possible
s	1
p	3
d	5
f	7

Here's an important fact: only two electrons can fit in each orbital. So, in an s orbital you can have a maximum of 2 electrons; in a d orbital you can have a maximum of 10 electrons; in any orbital there can only be two electrons.

Since a d sublevel has 5 orbitals (and each orbital can contain up to two electrons) then a d sublevel can contain 10 electrons ( $= 5 \times 2$ ). Pay attention to the difference between "sublevel" and "orbital".

## Critical Thinking Questions

- How many orbitals are there in a p sublevel? 3
- How many orbitals are there in a d sublevel? 5
- a) How many total sublevels would be found in the entire 2<sup>nd</sup> energy level? Two, since the 2<sup>nd</sup> energy level has s and p sublevels.  
b) How many orbitals would be found in the entire 2<sup>nd</sup> energy level? 4 orbitals: the 2s sublevel has 1 orbital and the 2p sublevel has 3 orbitals for a total of 4 orbitals.
- a) How many electrons can fit in an f sublevel?  
(2 electrons per orbital)  $\times$  (7 orbitals in an f sublevel) = 14 electrons.  
b) How many electrons can fit in an f orbital?  
Any orbital can only hold 2 electrons.
- How many electrons can fit in a d orbital? in a p orbital? in any kind of orbital?  
Each orbital can hold a maximum of 2 electrons.
- In your own words, what is the difference between a sublevel and an orbital?  
Orbitals are contained within sublevels. A sublevel is a grouping of orbitals.
- How many electrons can fit in each of the following energy levels:
  - 1<sup>st</sup> energy level = 2 because only 2 can fit in an s sublevel and the first energy level only has an s sublevel.
  - 2<sup>nd</sup> energy level = 8: 2 in the 2s sublevel and 6 in the 2p sublevel giving a total of 8.
  - 3<sup>rd</sup> energy level = 18: 2 in the 3s sublevel, 6 in the 3p sublevel, and 10 in the 3d sublevel giving a total of 18.
  - 4<sup>th</sup> energy level = 32: 2 in the 4s, 6 in the 4p, 10 in the 4d, and 14 in the 4f.

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