

# AP Chemistry Final Study Guide

## Unit 1: Atomic Structure and Properties (7%–9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
  - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
  - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
  - The greater the charge of the nucleus, the more energy is required to ionize an electron
  - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
  - When atoms absorb energy, electrons jump to higher energy levels
  - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
  - The faster an ejected electron is going, the more kinetic energy it has
  - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
  - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
    - Nickel:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$  (notice the d-block)
    - Nickel (with shorthand):  $[\text{Ar}] 4s^2 3d^8$
- **Law of Definite Proportions:**
  - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
  - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
  - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

# Chemistry Unit Study Guide

**Sasha Hu**



## Chemistry Unit Study Guide:

*PSI Study Guide for Organic Chemistry* by Douglas C. Neckers, Michael P. Doyle E. C. Blossey, 1977      *Ccea as Chemistry Student Unit Guide* Alyn McFarland, 2012 Written by a senior examiner Alyn G McFarland this CCEA AS Chemistry Student Unit Guide is the essential study companion for Unit 1 Basic Concepts in Physical and Inorganic Chemistry This full colour book includes all you need to know to prepare for your unit exam a aaa clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index a aaa examiner s advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required aa aaa exam style questions with graded student responses so you can see clearly what is required to get a better grade      **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office, 1968 Includes Part 1 Number 1 Books and Pamphlets Including Serials and Contributions to Periodicals January June      Introductory College Chemistry Neil Elbridge Gordon, 1926      *Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition* Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines      **Introductory College Chemistry** Harry Nicholls Holmes, 1926      Catalog of Copyright Entries, Third Series Library of Congress. Copyright Office, 1975 The record of each copyright registration listed in the Catalog includes a description of the work copyrighted and data relating to the copyright claim the name of the copyright claimant as given in the application for registration the copyright date the copyright registration number etc      **Introductory Chemistry** Neil Elbridge Gordon, 1927      *McGraw-Hill encyclopedia of science & technology* McGraw-Hill, 2002      **The United States Catalog** Mary Burnham, Carol Hurd, 1928      **The Indiana Teacher** , 1936      *Study Guide & Solutions Manual to Accompany Organic Chemistry, Third Edition* G. Marc Loudon, Joseph G. Stowell, 1995      Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 Barron s Science 360 Chemistry is your complete go to guide for everything chemistry This comprehensive guide is an essential resource for High school and college courses Homeschooling Virtual Learning Learning pods Inside you ll find Comprehensive Content Review Begin your study with the basic building block of chemistry and build as you go Topics include atomic structure chemical formulas electrochemistry the basics of organic chemistry and much more Effective Organization Topic organization and simple lesson formats break down the subject matter into manageable learning modules that help guide a successful study plan customized to your needs Clear Examples and Illustrations Easy to follow explanations hundreds of helpful illustrations and numerous step by step examples make this book ideal for self study and rapid learning Practice

Exercises Each chapter ends with practice exercises designed to reinforce and extend key skills and concepts These checkup exercises along with the answers and solutions will help you assess your understanding and monitor your progress Access to Online Practice Take your learning online for 50 practice questions designed to test your knowledge with automated scoring to show you how far you have come

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