

AP Chemistry Final Study Guide

Unit 1: Atomic Structure and Properties (7%-9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
 - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
 - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
 - The greater the charge of the nucleus, the more energy is required to ionize an electron
 - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
 - When atoms absorb energy, electrons jump to higher energy levels
 - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
 - The faster an ejected electron is going, the more kinetic energy it has
 - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
 - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
 - Nickel: $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$ (notice the d-block)
 - Nickel (with shorthand): $[\text{Ar}]4s^2 3d^8$
- **Law of Definite Proportions:**
 - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
 - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
 - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

Chemistry Unit Study Guide

Shasha Hu

Chemistry Unit Study Guide:

PSI Study Guide for Organic Chemistry by Douglas C. Neckers, Michael P. Doyle E. C. Blossey, 1977 Ccea as Chemistry Student Unit Guide Alyn McFarland, 2012 Written by a senior examiner Alyn G McFarland this CCEA AS Chemistry Student Unit Guide is the essential study companion for Unit 1 Basic Concepts in Physical and Inorganic Chemistry This full colour book includes all you need to know to prepare for your unit exam aaaa clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index aaaa examiner's advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required aaaa exam style questions with graded student responses so you can see clearly what is required to get a better grade

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