

Phase Change

- Deposition-gas to a solid; Sublimation-solid to gas; Vaporization- liquid to gas; Condensation-gas to liquid
- Physical change - melting, freezing, dissolving, evaporating;
- Chemical change- boiling an egg, converting, exploding, burning, digesting.

Atomic Numbers

- Number that identifies an element-Atomic number
- Weighted average mass of all occurring isotopes of an element-Atomic mass
- Symbolic method for expressing the composition of an atomic nucleus-Atomic notation

Electron Configurations

- Shorthand description of electrons by sublevels according to increasing energy-Electron configuration
- Sublevels: 2,6,10,14 electrons in the sublevels; 1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p;



Chemical Formulas

- Name of a chemical compound that indicates the # of atoms of each element-Chemical formula
- Number of atoms from formula- $\text{C}_7\text{H}_{16}\text{NO}_3$ = 40 atoms

NaHCO_3 -sodium hydrogen carbonate; $\text{Mg}(\text{ClO}_4)_2$ - Magnesium Perchlorate; PbO_2 - Lead (IV) oxide;
 Hg_2Br_2 -Mercury (II) bromide; $\text{Al}_2(\text{CrO}_4)_3$ -Aluminum Cromate; ZnCl_2 - Zinc Chloride; Nitrous acid- HNO_2 ;
Hypochlorous acid- HClO ; hydrochloric- HCl ; H_2S -hydrosulfuric acid; HClO_4 -perchloric acid; H_2SO_4 -sulfuric acid; HNO_3 -nitric acid; SnO_2 -tin dioxide; sulfuric acid- H_2SO_4 ;

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[Edition Donald W. Shive,Louise E. Shive,1986](#) [Study Guide, Introductory Chemistry, a Foundation ; Introductory](#)

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