

## Thermodynamics

Irreversible  $\rightarrow$  Independence of mass

Temp, Pressure, density, molar nc, specific nc, viscosity, MP/EP, surface viscosity

NOTE -

$$\frac{Ent}{Ext} = Inv$$

Extensive  $\rightarrow$   $g = Vol^m$ , heat cap, U, H, G, S  $du = dQ + dH$

Irreversible



- Process is fast
- one step

CONST. EXT. PRESSURE

Exch = +ve

Exch = -ve

W by system = -ve

W by system = +ve

Heat cap = heat reqd to raise temp by 1°C.

Heat cap  $\rightarrow$  param for  $\gamma$

molar heat cap/cp  $\rightarrow$   $\gamma$   $\rightarrow$   $\frac{1}{molar heat cap}$

$$C = \frac{Q}{\Delta T} \quad \text{for molar nc}$$

for specific nc  $\rightarrow$

$$C = \frac{Q}{(m) \Delta T}$$

$$C_m = \frac{Q}{m \Delta T}$$

Reversible adiabatic process

$$\rightarrow P V^{\gamma} = \text{const}$$

Diatomic linear  $\rightarrow f = 3+2 = 5$

Triatomic non linear  $\rightarrow f = 3+3 = 6$

Process	Work done	q	$\Delta U$	$\Delta H$
Rev. Isothermal	$-nRT \ln(V_2/V_1)$	-w	0	0
Irrev. Isenthalpic	$-P \Delta V (V_2 - V_1)$	-w	0	0

$$\Delta H = \text{enthalpy change} = nCp\Delta T$$

$$\Delta U = nCv \Delta T \rightarrow \text{true for all}$$

for Isobaric process,

$$\Delta H = q = nCp\Delta T$$

# Chemistry Thermodynamics Iit Jee Notes

**Galv**

## **Chemistry Thermodynamics Iit Jee Notes:**

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**Heat & Thermodynamics for JEE Main & Advanced (Study Package for Physics)** D. C. Er. Gupta,2016-12

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