

# Thermodynamics

Intensive  $\rightarrow$  Independent of mass

Temp, Pressure, density, molar HC, specific HC, molarity, M.P/B.P, surface viscosity

NOTE -

$$\frac{E_{\text{out}}}{E_{\text{in}}} = \text{Inc}$$

Extensive  $\rightarrow$  eg - vol<sup>n</sup>, heat cap, U, H, q, S /  $dU = dQ + dW$

Irreversible (Quasi P)



Process is fast one step

CONST. EXT. PRESSURE

Q<sub>sys</sub> = +ve

Q<sub>surround</sub> = -ve

W<sub>by system</sub> = -ve

W<sub>on system</sub> = +ve

Heat = heat reqd to raise temp by 1°C

Heat  $q_p \rightarrow$  path function

molar heat cap /  $c_p$

convert

$$C = \frac{Q}{\Delta T}$$

for molar HC

for specific HC  $\rightarrow$   
 $C = \frac{Q}{(m)\Delta T}$

$$C_m = \frac{Q}{n\Delta T}$$

Reversible adiabatic process

$$PV^\gamma = \text{const}$$

Monatomic linear  $\rightarrow f = 3 + 2 = 5$

Diatomic non linear  $\rightarrow f = 3 + 3 = 6$

Process

W<sub>done</sub>

q

$\Delta U$

$\Delta H$

Rev. Isothermal

$$-nRT \ln(V_2/V_1)$$

-w

0

0

Irrev. Isothermal

$$-P_{\text{ext}}(V_2 - V_1)$$

-w

0

0

$$\Delta H = \text{enthalpy change} = nC_p \Delta T$$

for isobaric process,

$$\Delta H = q = nC_p \Delta T$$

$$\Delta U = nC_v \Delta T$$

True for all

# Chemistry Thermodynamics Iit Jee Notes

**Galv**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

## **Chemistry Thermodynamics IIT JEE Notes:**

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