

CHEMISTRY Terminology

The ABCs of chemistry

A

A: Symbol for atomic mass number
absolute zero: Lowest temperature; 0 Kelvin
Ae: Symbol for the element actinium, $Z = 89$
accuracy: The closeness of measurement to true value
acetate anion: $C_2H_3O_2^-$, charge -1
acetic acid: $C_2H_4O_2$
acetone: Dimethyl ketone; organic solvent
acid-base indicator: A chemical that changes color if the pH changes
acid-base titration: A method for determining acid or base concentrations
acid definition: See Arrhenius acid, Brønsted-Lowry acid, Lewis acid
acid ionization constant (K_a): Equilibrium constant for acid dissociation
acidic solution: pH below 7
actinide: Element with $Z = 90$ or above; radioactive
activation energy (E_a): A process's energy barrier



adhesion: Attraction of unlike molecules or materials
Ag: Symbol for the element silver, $Z = 47$
Al: Symbol for the element aluminum, $Z = 13$
alcohol: Organic compound with -OH group; ROH
aldehyde: Carbonyl group bonded to hydrogen and an organic group
alkali metal family: Lithium, sodium, potassium, rubidium, cesium, and francium; column #1
alkaline earth metal family: Beryllium, magnesium, calcium, strontium, barium, and radium; column #2
alkane: Hydrocarbon; all C-C single bonds
alkene: Hydrocarbon; 1 or more C=C double bonds
alkyne: Hydrocarbon; 1 or more C≡C triple bonds
allotropes: Two or more forms of an element
alloy: Solution of 2 or more metals
alpha (α): Greek letter that denotes radioactive particle and various scientific variables
alpha particle: Helium nucleus; charge +2
Au: Symbol for the element americium, $Z = 95$
amalgam: Alloy of mercury with other metal
amine: Organic base; RRRN; modified ammonia
amino acid: Compound with organic acid and organic base property; forms proteins and peptides
ammonia: NH_3 , base
ammonium cation: NH_4^+ , charge +1
ammonium hydroxide: NH_4OH
amphoteric oxide: Exhibits both acid and basic properties
amu: Atomic mass unit; $1/12$ mass of C-12
anion: Ion with a negative charge
anode: Electrode that supports oxidation
antibonding MO: MO is less stable than separate AOs
AO: Abbreviation for atomic orbital

aq: Aqueous; "dissolved in water"
Ar: Symbol for the element argon, $Z = 18$
aromatic: Organic compound with a benzene ring
Arrhenius acid: Produces hydronium ion in water solution
Arrhenius base: Yields hydroxide ion in water solution
Arrhenius equation for rate constant (k):

$$k = Ae^{-E_a/RT}$$
 E_a = activation energy
arsenide anion: As^{3-} , charge -3
As: Symbol for the element arsenic, $Z = 33$
At: Symbol for the element astatine, $Z = 85$
atm: Symbol for pressure in "atmospheres"
atom: Fundamental unit of all matter
atomic mass number (A): Total number of protons and neutrons
atomic number (Z): Number of protons in the nucleus
atomic orbital: Wave motion of electrons in atoms
atomic radius: Empirical measure of atom size
atomic weight: Weighted average of natural isotopes of an element
Au: Symbol for the element gold, $Z = 79$
Aufbau principle: Guides the filling of electronic subshells of the elements
Avogadro's law: Volume is proportional to gas moles at fixed pressure and temperature
Avogadro's number: $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

B

B: Symbol for the element boron, $Z = 5$
Ba: Symbol for the element barium, $Z = 56$
balanced equation: Each side of the equation has the same number of atoms of each element and charge
balanced redox: Electron loss = electron gain
barium cation: Ba^{2+} , charge +2
base definition: See Arrhenius base, Brønsted-Lowry base, Lewis base
base ionization constant (K_b): Equilibrium constant for base dissociation
basic solution: pH over 7
battery: Source of electrical power; galvanic cell
Be: Symbol for the element beryllium, $Z = 4$
benzene: C_6H_6 ; organic solvent; aromatic ring of 6 carbon atoms
beryllium cation: Be^{2+} , charge +2
beta (β): Greek letter that denotes radioactive particle and a number of scientific variables
beta particle: Energetic electron from the nucleus; charge -1
Bh: Symbol for the element bohrium, $Z = 107$
Bi: Symbol for the element bismuth, $Z = 83$
bicarbonate anion: HCO_3^- , charge -1
bimolecular oxygen: O_2
binary compound: Formed from 2 elements
bisulfate anion: HSO_4^- , charge -1



Bk: Symbol for the element berkelium, $Z = 97$
bleach: Chemical that is a strong oxidizing agent
boiling: Liquid → gas at the boiling point
boiling point elevation: A solution has a higher boiling point than a pure solvent
boiling point (T_b): Liquid-gas equilibrium, $P = 1$ atm
boiling point of water: 100°C
bond energy: Energy held by a chemical bond
bond length: Distance between 2 atoms in a bond
bond order for Lewis model: # of bonds divided by # of bonded atoms
bond order for MO treatment: # of filled bonding MOs minus # of filled antibonding MOs
bonding electrons: Form chemical bonds in a compound
bonding MO: MO is more stable than AOs
Boyle's law: PV constant for gas, fixed T and n
Br: Symbol for the element bromine, $Z = 35$
bromide anion: Br^- , charge -1
Brønsted-Lowry acid: Proton donor
Brønsted-Lowry base: Proton acceptor
buffer: Composed of weak acid and weak base; serves to keep pH constant



C

c: Symbol for c atch, the SI prefix for $1/100$ or 10^{-2}
c: Symbol for the speed of light
°C: Celsius temperature scale
C: Symbol for the element carbon, $Z = 6$
 $C_2H_4O_2$: Acetic acid
Ca: Symbol for the element calcium, $Z = 20$
calcium cation: Ca^{2+} , charge +2
calcium fluoride: CaF_2
calcium hydroxide: $Ca(OH)_2$
calorie: English unit of heat energy
calorimetry: Study of the release or absorption of heat
carbide anion: C^{4-} , charge -4
carbohydrate: Organic compound; carbon bonded to several -OH groups (sugar and starch)
carbon dioxide: CO_2
carbon monoxide: CO
carbonate anion: CO_3^{2-} , charge -2
carbonic acid: H_2CO_3
carbonyl group: CO , ligand complex
carbonyl group: $>CO$ bonded to 2 groups
carboxylic acid: Organic acid; $R-COOH$
catalyst: Accelerates a reaction but is not consumed in the reaction; lowers the activation energy
cathode: Electrode that supports reduction
cation: Ion with a positive charge
Cd: Symbol for the element cadmium, $Z = 48$



Chemistry Vocabulary Study Guide

Lewis Morris



Chemistry Vocabulary Study Guide:

Chemistry Regents Vocabulary Workbook Lewis Morris, Learn the Secret to Success on the Chemistry Regents Exam Ever wonder why learning comes so easily to some people This remarkable workbook reveals a system that shows you how to learn faster easier and without frustration By mastering the hidden language of the subject and exams you will be poised to tackle the toughest of questions with ease We ve discovered that the key to success on the Chemistry Regents Exam lies with mastering the Insider s Language of the subject People who score high on their exams have a strong working vocabulary in the subject tested They know how to decode the vocabulary of the subject and use this as a model for test success People with a strong Insider s Language consistently Perform better on their Exams Learn faster and retain more information Feel more confident in their courses Perform better in upper level courses Gain more satisfaction in learning The Chemistry Regents Exam Vocabulary Workbook is different from traditional review books because it focuses on the exam s Insider s Language It is an outstanding supplement to a traditional review program It helps your preparation for the exam become easier and more efficient The strategies puzzles and questions give you enough exposure to the Insider Language to use it with confidence and make it part of your long term memory The Chemistry Regents Exam Vocabulary Workbook is an awesome tool to use before a course of study as it will help you develop a strong working Insider s Language before you even begin your review Learn the Secret to Success After nearly 20 years of teaching Lewis Morris discovered a startling fact Most students didn t struggle with the subject they struggled with the language It was never about brains or ability His students simply didn t have the knowledge of the specific language needed to succeed Through experimentation and research he discovered that for any subject there was a list of essential words that when mastered unlocked a student s ability to progress in the subject Lewis called this set of vocabulary the Insider s Words When he applied these Insider s Words the results were incredible His students began to learn with ease He was on his way to developing the landmark series of workbooks and applications to teach this Insider s Language to students around the world

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The Routledge Handbook of Vocabulary Studies Stuart Webb,2019-07-30 The Routledge Handbook of Vocabulary Studies provides a cutting edge survey of current scholarship in this area Divided into four sections which cover understanding vocabulary approaches to teaching and learning vocabulary measuring knowledge of vocabulary and key issues in teaching researching and measuring vocabulary this Handbook brings together a wide range of approaches to learning words to provide clarity on how best vocabulary might be taught and learned provides a comprehensive discussion of the key issues and challenges in vocabulary studies with research taken from the past 40 years includes chapters on both formulaic language as well as single word items features original contributions from a range of internationally renowned scholars as well as academics at the forefront of innovative research The Routledge Handbook of Vocabulary Studies is an essential text for those interested in teaching learning and researching vocabulary

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series of workbooks and applications to teach this Insider's Language to students around the world **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 Guide to the Study and Use of Reference Books, 1908 **Guide to the Study and Use of Reference Books** Alice Bertha Kroeger, 1904 *AP Chemistry Exam Success* Lewis Morris, 2018-07-14

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A thesaurus of medical words and phrases, by W.M. Barton and W.A. Wells Wilfred Mason Barton,1903

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
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