

CHAPTER 15 STUDY GUIDE

Energy and Chemical Change

Section 15.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- _____ 1. *Energy* is the ability to do work or produce heat.
- _____ 2. The law of conservation of energy states that energy *can be* created and destroyed.
- _____ 3. Chemical potential energy is energy stored in a substance because of its *composition*.
- _____ 4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
- _____ 5. A calorie is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
- _____ 6. A *calorie* is the SI unit of heat and energy.
- _____ 7. The *specific* heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.
- _____ 8. *Kinetic energy* is energy of motion.
- _____ 9. Chemicals participating in a chemical reaction contain only *potential energy*.
- _____ 10. One nutritional Calorie is equal to *100 calories*.
- _____ 11. One calorie equals *4.184 joules*.
- _____ 12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
- _____ 13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

Chemistry Study Guide Energy And Chemical Change

J.A. (Sandy) Sprinkle



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Basic Concepts of Chemistry, Study Guide Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems

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Table of Contents **Chemistry Study Guide Energy And Chemical Change**

1. Understanding the eBook **Chemistry Study Guide Energy And Chemical Change**
 - The Rise of Digital Reading **Chemistry Study Guide Energy And Chemical Change**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Chemistry Study Guide Energy And Chemical Change**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Chemistry Study Guide Energy And Chemical Change**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Chemistry Study Guide Energy And Chemical Change**
 - Personalized Recommendations
 - **Chemistry Study Guide Energy And Chemical Change** User Reviews and Ratings

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- 5. Accessing Chemistry Study Guide Energy And Chemical Change Free and Paid eBooks
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- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Chemistry Study Guide Energy And Chemical Change
 - Interactive Elements Chemistry Study Guide Energy And Chemical Change
- 8. Staying Engaged with Chemistry Study Guide Energy And Chemical Change
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Energy And Chemical Change
- 9. Balancing eBooks and Physical Books Chemistry Study Guide Energy And Chemical Change
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Energy And Chemical Change
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Study Guide Energy And Chemical Change
 - Setting Reading Goals Chemistry Study Guide Energy And Chemical Change
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Study Guide Energy And Chemical Change
 - Fact-Checking eBook Content of Chemistry Study Guide Energy And Chemical Change
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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