

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



atom

J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

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Mary H. Scanlan



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Table of Contents Chemistry Study Guide Hawthorne High School

1. Understanding the eBook Chemistry Study Guide Hawthorne High School
 - The Rise of Digital Reading Chemistry Study Guide Hawthorne High School
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Hawthorne High School
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Hawthorne High School
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Hawthorne High School
 - Personalized Recommendations
 - Chemistry Study Guide Hawthorne High School User Reviews and Ratings
 - Chemistry Study Guide Hawthorne High School and Bestseller Lists

5. Accessing Chemistry Study Guide Hawthorne High School Free and Paid eBooks
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 - Chemistry Study Guide Hawthorne High School Budget-Friendly Options
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 - Chemistry Study Guide Hawthorne High School Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide Hawthorne High School
 - Highlighting and Note-Taking Chemistry Study Guide Hawthorne High School
 - Interactive Elements Chemistry Study Guide Hawthorne High School
8. Staying Engaged with Chemistry Study Guide Hawthorne High School
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Hawthorne High School
9. Balancing eBooks and Physical Books Chemistry Study Guide Hawthorne High School
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Hawthorne High School
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Study Guide Hawthorne High School
 - Setting Reading Goals Chemistry Study Guide Hawthorne High School
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide Hawthorne High School
 - Fact-Checking eBook Content of Chemistry Study Guide Hawthorne High School
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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