

Thermochemistry HOMEWORK

Lesson 1 - Homework

1. Imagine a bowl of soup in the kitchen. The soup is the system. Describe the surroundings.

The surroundings include the bowl, the air, and the entire room except it.

2. Complete the sentence with the words endothermic/exothermic; and positive/negative. When heat is absorbed by the system from the surroundings, that process is said to be endothermic and the sign of q is positive.

3. Which is not true for an endothermic reaction?

- The temperature of the surroundings decreases.
- The heat for the reaction is positive.
- Heat flows from the surroundings into the system.
- The products have a lower heat energy than the reactants.
- All of the above are true.

4. If the same amount of heat is added to 5.00 g samples of each of the metals below, which metal will experience the greatest temperature change?

| Metal | Specific Heat Capacity (J/g \cdot °C) |
|-----------|---|
| Al | 0.907 |
| <u>Sn</u> | <u>0.205</u> |
| Cu | 0.385 |
| Fe | 0.449 |
| Pb | 0.129 |

The metal with the LOWEST specific heat capacity will require the LEAST amount of energy to increase its temperature.

If the same amount of heat is added to each metal, the one with the LOWEST specific heat capacity will reach the highest temp.

5. Sometimes, specific heat capacities are given in molar instead of grams, with the units J/mol \cdot °C. If the molar heat capacity of a compound with the formula C_2H_6O is 88 J/mol \cdot °C, what is the specific heat (in J/g \cdot °C) of this substance?

$$88 \text{ J/mol}^\circ\text{C} \times (1 \text{ mol} / 46.07 \text{ g}) = 1.9103537 \text{ J/g}^\circ\text{C} \rightarrow 1.91 \text{ J/g}^\circ\text{C}$$

6. A sample of aluminum (0.90 J/g \cdot °C) absorbs 4.32 J of energy while being heated from 23.2 °C to 38.5 °C. Determine the mass of the sample.

$$\text{J/g}^\circ\text{C} = (\text{mass} \text{ (J/g}^\circ\text{C)}) \times (\Delta T)$$

$$0.90 \text{ J/g}^\circ\text{C} = (x \text{ g}) (38.5 - 23.2)^\circ\text{C}$$

$$x = 0.071142857 \text{ g} \rightarrow 0.071 \text{ g}$$

It's necessary to convert the temperatures to Celsius. This is because the CHANGE in Celsius is the same as the CHANGE in Kelvin. Since you're given both temperatures, you can find the CHANGE in temperature. If you chose using just temperature, you would need to convert the single temperature into Celsius.

7. An iron sphere with a volume of 3.30 m³ and density of 7.86 g/cm³ at 25.0 °C is dropped into a perfectly insulated bucket containing 50.0 g of H₂O at 100.0 °C. The specific heats of Fe and H₂O are 0.450 and 4.184 J/g \cdot °C, respectively. What is the final temperature (in °C) of the iron sphere?

$$\text{The volume and density is first given:}$$

$$3.30 \text{ m}^3 \times (7.86 \text{ g/cm}^3) = 26.058 \text{ kg}$$

$$\text{Heat gained} = \text{Heat lost}$$

$$-q_{\text{Fe}} = q_{\text{H}_2\text{O}}$$

$$-(\text{mass}_{\text{Fe}})(C_p)(T_f - T_i) = (\text{mass}_{\text{H}_2\text{O}})(C_p)(T_f - T_i)$$

$$26.058 \text{ kg} = 26058 \text{ g}$$

$$-(26058 \text{ g})(0.450 \text{ J/g}^\circ\text{C})(T_f - 25.0^\circ\text{C}) = (50.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(T_f - 100.0^\circ\text{C})$$

$$-11726.1 \text{ J} - 11726.1 \text{ J} = -170 \text{ J} T_f = 14549.8 \text{ J}$$

$$14549.8 \text{ J} = 14549.8 \text{ J}$$

$$T_f = 171.882388^\circ\text{C} \rightarrow 172^\circ\text{C} - 273 = 44^\circ\text{C}$$

The lower the number with LOWEST the heat increase it is meaning it is higher temperature.

8. Given that silver has a specific heat of 0.235 J/g \cdot °C.

a. If a 75.0 g sample of silver at 100. °C is dropped into a calorimeter (perfectly insulated container) containing 100. g of water at 22.0 °C, what will be the final temperature of the system at equilibrium? (The specific heat capacity of water is 4.184 J/g \cdot °C).

$$\text{Heat gained} = \text{Heat lost}$$

$$-q_{\text{Ag}} = q_{\text{H}_2\text{O}}$$

$$-(\text{mass}_{\text{Ag}})(C_p)(T_f - T_i) = (\text{mass}_{\text{H}_2\text{O}})(C_p)(T_f - T_i)$$

$$75.0 \text{ g} = 75.0 \text{ g}$$

$$-(75.0 \text{ g})(0.235 \text{ J/g}^\circ\text{C})(T_f - 100.0^\circ\text{C}) = (100. \text{ g})(4.184 \text{ J/g}^\circ\text{C})(T_f - 22.0^\circ\text{C})$$

$$-17.625 \text{ J} - 17.625 \text{ J} = -418.4 \text{ J} T_f = 418.4 \text{ J}$$

$$418.4 \text{ J} = 418.4 \text{ J}$$

$$T_f = 246.222789^\circ\text{C} \rightarrow 246^\circ\text{C} - 273 = 73^\circ\text{C}$$

b. How much heat energy is transferred to the water in part a?

$$\text{J/g}^\circ\text{C} \times (\text{mass} \text{ (J/g}^\circ\text{C)}) \times (\Delta T)$$

$$4.184 \text{ J/g}^\circ\text{C} \times 100. \text{ g} \times (73 - 22)^\circ\text{C}$$

$$21111.2 \text{ J} \rightarrow 21111 \text{ J}$$

9. A 75.0 g sample of water at 80.0 °C is added to an unknown amount of water at 22.0 °C in a perfectly insulated container. The final temperature is 50.0 °C. How many grams of water were there initially at 22.0 °C? (The specific heat capacity of water is 4.184 J/g \cdot °C).

$$\text{Heat gained} = \text{Heat lost}$$

$$-q_{\text{H}_2\text{O}} = q_{\text{H}_2\text{O}}$$

$$-(\text{mass}_{\text{H}_2\text{O}})(C_p)(T_f - T_i) = (\text{mass}_{\text{H}_2\text{O}})(C_p)(T_f - T_i)$$

$$75.0 \text{ g} = 75.0 \text{ g}$$

$$-(75.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(50.0^\circ\text{C} - 80.0^\circ\text{C}) = (x \text{ g})(4.184 \text{ J/g}^\circ\text{C})(50.0^\circ\text{C} - 22.0^\circ\text{C})$$

$$-10665 \text{ J} = -10665 \text{ J}$$

$$x = 100.000000 \text{ g} \rightarrow 100 \text{ g}$$

10. The specific heat capacity of copper is 0.385 J/g \cdot °C. What quantity of heat is required to heat 286 g of copper from -6.0 °C to +12.0 °C? (The specific heat capacity of water is 4.184 J/g \cdot °C).

$$\text{J/g}^\circ\text{C} \times (\text{mass} \text{ (J/g}^\circ\text{C)}) \times (\Delta T)$$

$$0.385 \text{ J/g}^\circ\text{C} \times 286 \text{ g} \times (12 - (-6))^\circ\text{C}$$

$$3904.7 \text{ J} \rightarrow 3905 \text{ J}$$

Chemistry Thermochemistry Homework Packet Answers

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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

Learning about Materials Ted Lister, Colin Osborne, 1998 This book presents useful material in the form of teacher's notes and photocopiable worksheets It will be useful at pre and post 16 level

Synthesis and Characterization of Nitric Oxide-releasing Agents/polymers for Biomedical Applications Melissa May

Batchelor, 2004

Publications of the National Bureau of Standards, 1979 Catalog United States. National Bureau of Standards, 1980

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High School Chemistry Teachers Magazine, 1967

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Photodissociation of Ketene Melonie Anne Teichert, 1999

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The British National Bibliography Arthur James Wells, 1979

Homework Questions and Answers for Freshman Chemistry Loren Sheldon Carter, Ellis Ray Peterson, 1987

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Chemistry Homework Frank Schaffer Publications, Joan DiStasio, 1996-03 Includes the periodic table writing formulas balancing equations stoichiometry problems and more

Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble, 1994

Chemistry 270: Clicker Questions Christopher Nichols, 2013-07-29

Student Workbook and Selected Solutions Manual for Introductory Chemistry Steve Russo, Michael Silver, 2014-02-07 Features examples from each chapter learning objectives review of key concepts from the text and additional problems for student practice Also the workbook provides comprehensive answers and explanations to selected end of chapter problems from the text Provides over 200 worked examples and more than 550 practice problems and quiz questions to help students develop and practice their problem solving skills

A Level Chemistry Multiple Choice Questions and Answers (MCQs) Arshad Iqbal, 2020-04-10

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Table of Contents Chemistry Thermochemistry Homework Packet Answers

1. Understanding the eBook Chemistry Thermochemistry Homework Packet Answers
 - The Rise of Digital Reading Chemistry Thermochemistry Homework Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Thermochemistry Homework Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Thermochemistry Homework Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Thermochemistry Homework Packet Answers
 - Personalized Recommendations
 - Chemistry Thermochemistry Homework Packet Answers User Reviews and Ratings
 - Chemistry Thermochemistry Homework Packet Answers and Bestseller Lists
5. Accessing Chemistry Thermochemistry Homework Packet Answers Free and Paid eBooks
 - Chemistry Thermochemistry Homework Packet Answers Public Domain eBooks
 - Chemistry Thermochemistry Homework Packet Answers eBook Subscription Services
 - Chemistry Thermochemistry Homework Packet Answers Budget-Friendly Options

6. Navigating Chemistry Thermochemistry Homework Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Thermochemistry Homework Packet Answers Compatibility with Devices
 - Chemistry Thermochemistry Homework Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Thermochemistry Homework Packet Answers
 - Highlighting and Note-Taking Chemistry Thermochemistry Homework Packet Answers
 - Interactive Elements Chemistry Thermochemistry Homework Packet Answers
8. Staying Engaged with Chemistry Thermochemistry Homework Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Thermochemistry Homework Packet Answers
9. Balancing eBooks and Physical Books Chemistry Thermochemistry Homework Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Thermochemistry Homework Packet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Thermochemistry Homework Packet Answers
 - Setting Reading Goals Chemistry Thermochemistry Homework Packet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Thermochemistry Homework Packet Answers
 - Fact-Checking eBook Content of Chemistry Thermochemistry Homework Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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