

Moles and

Reactions

Information: Mole Ratios in Equations

Propane is burned in many rural homes for heat in the winter. Below is the balanced equation for the combustion of propane (C₃H₈).



For each molecule of propane that is burned, there needs to be five molecules of oxygen present. Likewise, if there were a dozen molecules of propane, five dozen molecules of oxygen would be required. Similarly, for each mole of propane, five moles of oxygen are needed. Also, for each mole of propane burned three moles of carbon dioxide and 4 moles of water are produced. The numbers of moles of each substance in a chemical equation are related by the ratio of the coefficients of each substance.

Critical Thinking Questions

Note: For questions 1-6, refer to the balanced equation for the combustion of propane.

1. a) How many moles of water are produced when 1.45 moles of propane are combusted?
It is a 4:1 ratio so the moles of water = $(1.45)(4) = 5.8$ moles
1. b) How many molecules of water is this? (Remember each mole has 6.02×10^{23} molecules.)
 $(5.8)(6.02 \times 10^{23}) = 3.49 \times 10^{24}$
2. If 2.35 moles of CO₂ are produced in a reaction, how many moles of H₂O would be produced?
It is a 4:3 ratio so the moles of water = $2.35(4/3) = 3.13$ moles
3. Why is this statement false: "If 10 grams of propane burn, you need 50 grams of oxygen." The coefficients in the balanced equation relate mole ratios and not gram ratios. The above statement would be true if it read: "If 10 moles of propane burn, you need 50 moles of oxygen."
4. a) If 27.3 moles of carbon dioxide are produced during the combustion of a certain amount of propane, how many moles of propane were combusted?
Since it is a 1:3 ratio, the moles of propane = $27.3(1/3) = 9.1$ moles
4. b) How many grams of propane was this?

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David Kirk

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