



Time (s)	Moles N ₂ O ₅	Moles NO ₂	Moles O ₂
0	0.4120	0	0
600	0.4194	0.0252	0.0063
1200	0.4104	0.0432	0.0108
1800	0.4032	0.0576	0.0144 \rightarrow 0.0036

$\underline{[2(0.0036)]} \quad \underline{[4(0.0036)]}$

Critical Thinking Questions

7. Calculate the change in moles of each reactant and product between 600 and 1200 seconds.

change in moles of N₂O₅: $2:4$ change in moles of NO₂: $4:1$ change in moles of O₂:
 -0.009 $+0.018$ $+0.0045$

8. Compare your answers in question 7. What is the relationship between the coefficients in the balanced chemical equation and the number of moles used up or produced in a reaction?

9. Fill in the missing blanks in the table above.

10. Calculate the following values. Be sure to use molarity in your calculations.

a) Rate of disappearance of N₂O₅ between time 0 and 600 s.

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CHEMICAL WEEK ,1996

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