



Time (s)	Moles N_2O_5	Moles NO_2	Moles O_2
0	0.4320	0	0
600	0.4194	0.0252	0.0063
1200	0.4104	0.0432	0.0108
1800	0.4032	0.0576	0.0144

$$2(0.0288) + 4(0.0036)$$

Critical Thinking Questions

7. Calculate the change in moles of each reactant and product between 600 and 1200 seconds.

change in moles of N_2O_5 : $2:4$ change in moles of NO_2 : $4:1$ change in moles of O_2 :

$= 0.009$ $+ 0.018$ $+ 0.0045$

8. Compare your answers in question 7. What is the relationship between the coefficients in the balanced chemical equation and the number of moles used up or produced in a reaction?
9. Fill in the missing blanks in the table above.
10. Calculate the following values. Be sure to use molarity in your calculations.
- a) Rate of disappearance of N_2O_5 between time 0 and 600 s.

Chemquest 42 Answers

Kennedy Information (Firm)



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Modern Healthcare ,1990
Information (Firm),2002

CHEMICAL WEEK ,1996

The Directory of Management Consultants, 2003 Kennedy

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